

## Acids & Bases Review

1. Name the following compounds as acids and circle which are weak acids:



nitric acid



sulfuric acid



hydrofluoric acid



carbonic acid



acetic acid



phosphoric acid

2. Write the formulas for the following bases:

a) calcium hydroxide



b) aluminum hydroxide



c) ammonium hydroxide



d) lithium hydroxide



3. Complete the following table:

	Acids	Bases
Definition	A substance that increases the concentration of $\text{H}^+$ (or $\text{H}_3\text{O}^+$ ) when added to water. or A substance that acts as a proton ( $\text{H}^+$ ) donor.	A substance that increases the concentration of $\text{OH}^-$ when added to water. or A substance that acts as a proton ( $\text{H}^+$ ) acceptor.
Electrolyte in aqueous solution?	$\text{H}^+$	$\text{OH}^-$
pH	less than 7	greater than 7
Taste	sour	bitter
Litmus turns	blue turns red	red turns blue
Which reacts with metals to form $\text{H}_2$ gas?	yes	no

**Conjugate acid-base pairs:**

4. Identify the acid, the base, the conjugate acid and the conjugate base in the reactions:



5. An acid will form the conjugate base and a base will form the conjugate acid.

# Holt Chemistry Acid And Bases Test Answers

**Eileen Ramsden**



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