

# Graphing Periodic Trends Lab

Name \_\_\_\_\_ Hr \_\_\_\_\_

**PURPOSE:** To determine how certain properties are periodic when the elements are arranged in periods or groups.

**DISCUSSION:** There are many periodic trends that occur as you cross a period or go down a group/family on the Periodic Table. Graphing various elemental properties will help to illustrate these trends.

## REFERENCE:

Angstrom - A unit of distance equal to one ten-billionth of a meter.

Atomic Radius - The distance from the center of the nucleus to the valence electron energy level.

Ionization energy - The energy required to remove an electron from an atom.

## DATA

**Properties of the First 18 Elements**

Element	Atomic Number	Atomic Radius (Angstroms)	Ionization Energy
H	1	0.79	314
He	2	0.49	567
Li	3	2.05	124
Be	4	1.4	215
B	5	1.17	191
C	6	0.91	260
N	7	0.75	335
O	8	0.65	314
F	9	0.57	402
Ne	10	?	?
Na	11	2.23	119
Mg	12	1.72	176
Al	13	1.82	138
Si	14	1.46	168
P	15	1.23	242
S	16	1.09	239
Cl	17	0.97	299
Ar	18	?	?

Using the data above, make a line graph the following data. Draw a dark vertical line to separate the 3 periods. Give each graph a title and label both the x-axis and the y-axis. Connect your points. Predict the plot points for elements with question marks.

1. Atomic Radius (y) vs. Atomic Number (x) for the first 3 periods (18 elements – 3 graphs)
2. Ionization energy (y) vs. Atomic Number (x) for the first 3 periods (18 elements – 3 graphs)

# Graphing Trends In The Periodic Table Lab Answers

**Sebastian Brünink**



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## **Table of Contents Graphing Trends In The Periodic Table Lab Answers**

1. Understanding the eBook Graphing Trends In The Periodic Table Lab Answers
  - The Rise of Digital Reading Graphing Trends In The Periodic Table Lab Answers
  - Advantages of eBooks Over Traditional Books
2. Identifying Graphing Trends In The Periodic Table Lab Answers
  - Exploring Different Genres
  - Considering Fiction vs. Non-Fiction
  - Determining Your Reading Goals
3. Choosing the Right eBook Platform
  - Popular eBook Platforms
  - Features to Look for in an Graphing Trends In The Periodic Table Lab Answers
  - User-Friendly Interface
4. Exploring eBook Recommendations from Graphing Trends In The Periodic Table Lab Answers
  - Personalized Recommendations
  - Graphing Trends In The Periodic Table Lab Answers User Reviews and Ratings
  - Graphing Trends In The Periodic Table Lab Answers and Bestseller Lists
5. Accessing Graphing Trends In The Periodic Table Lab Answers Free and Paid eBooks

- Graphing Trends In The Periodic Table Lab Answers Public Domain eBooks
  - Graphing Trends In The Periodic Table Lab Answers eBook Subscription Services
  - Graphing Trends In The Periodic Table Lab Answers Budget-Friendly Options
6. Navigating Graphing Trends In The Periodic Table Lab Answers eBook Formats
    - ePub, PDF, MOBI, and More
    - Graphing Trends In The Periodic Table Lab Answers Compatibility with Devices
    - Graphing Trends In The Periodic Table Lab Answers Enhanced eBook Features
  7. Enhancing Your Reading Experience
    - Adjustable Fonts and Text Sizes of Graphing Trends In The Periodic Table Lab Answers
    - Highlighting and Note-Taking Graphing Trends In The Periodic Table Lab Answers
    - Interactive Elements Graphing Trends In The Periodic Table Lab Answers
  8. Staying Engaged with Graphing Trends In The Periodic Table Lab Answers
    - Joining Online Reading Communities
    - Participating in Virtual Book Clubs
    - Following Authors and Publishers Graphing Trends In The Periodic Table Lab Answers
  9. Balancing eBooks and Physical Books Graphing Trends In The Periodic Table Lab Answers
    - Benefits of a Digital Library
    - Creating a Diverse Reading Collection Graphing Trends In The Periodic Table Lab Answers
  10. Overcoming Reading Challenges
    - Dealing with Digital Eye Strain
    - Minimizing Distractions
    - Managing Screen Time
  11. Cultivating a Reading Routine Graphing Trends In The Periodic Table Lab Answers
    - Setting Reading Goals Graphing Trends In The Periodic Table Lab Answers
    - Carving Out Dedicated Reading Time
  12. Sourcing Reliable Information of Graphing Trends In The Periodic Table Lab Answers
    - Fact-Checking eBook Content of Graphing Trends In The Periodic Table Lab Answers
    - Distinguishing Credible Sources
  13. Promoting Lifelong Learning
    - Utilizing eBooks for Skill Development

- Exploring Educational eBooks

### 14. Embracing eBook Trends

- Integration of Multimedia Elements
- Interactive and Gamified eBooks

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