

Grandma Mole's Fudge**Introduction**

Grandma Mole was a chemist and she kept all of her recipes in units of moles in order to keep them a family secret. In this lab you will need to convert her recipes with the chemical names into English units of measurement in order to be able to make and EAT the fudge.

Pre-Lab Questions – due when you walk into the lab FRIDAY.

1. What safety precautions should you take when performing this lab? (2 pt. each)

- Wear _____
- Wear _____
- Be careful around the _____
- Let the _____ cool before _____

Pre-Lab Data Calculations – due when you walk into the lab FRIDAY.

1. Sucrose is the disaccharide found in table sugar that you will use to sweeten your fudge. Find the molar mass of sucrose, $C_{12}H_{22}O_{11}$.

Use the given conversion factors below to find the recipe. Fill in the chart on the back with the actual measurements. (Show all work.)

2. 1.33 mol powdered sugar \rightarrow _____ lb powdered sugar

$$242.3 \text{ g powdered sugar} = 1 \text{ mole powdered sugar}$$

$$453.6 \text{ g powdered sugar} = 1 \text{ pound powdered sugar}$$

3. 0.890 mol cocoa powder \rightarrow _____ cup cocoa powder

$$1 \text{ mole cocoa powder} = 83.7 \text{ g cocoa powder}$$

$$325 \text{ g cocoa powder} = 1 \text{ cup cocoa powder}$$

4. 0.403 mol NaCl \rightarrow _____ teaspoon salt

$$1 \text{ mole NaCl} = 58.4 \text{ g NaCl}$$

$$2.350 \text{ g NaCl} = 1 \text{ mL salt}$$

$$1 \text{ mL salt} = 0.02 \text{ teaspoon salt}$$

Grandma Moles Fudge Lab Answers

BM King



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