

Calculating Specific Heat Worksheet

Name: _____

Date: _____



$Q = mc\Delta T$, where Q = heat energy, m = mass, and ΔT = change in temp. Remember, $\Delta T = (T_{\text{final}} - T_{\text{initial}})$. Show all work and proper units.

1. A 15.75-g piece of iron absorbs 1086.75 joules of heat energy, and its temperature changes from 25°C to 175°C. Calculate the specific heat capacity of iron.

$$Q = (m)(c)(\Delta T)$$

$$1086.75 = (15.75)(c)(175 - 25)$$

$$1086.75 = 2362.5c$$

$$\frac{1086.75}{2362.5} = \frac{2362.5c}{2362.5}$$

$$c = 0.4672^\circ\text{C}$$

2. How many joules of heat are needed to raise the temperature of 10.0 g of aluminum from 22°C to 55°C, if the specific heat of aluminum is 0.90 J/g°C?

$$Q = (m)(c)(\Delta T)$$

$$Q = (10)(0.90)(55 - 22)$$

$$Q = 297\text{ J}$$

3. Calculate the specific heat capacity of a piece of wood if 1500.0 g of the wood absorbs 67,500 joules of heat, and its temperature changes from 32°C to 57°C.

$$Q = (m)(c)(\Delta T)$$

$$67500 = (1500)(c)(57 - 32)$$

$$67500 = 37500c$$

$$\frac{67500}{37500} = \frac{37500c}{37500}$$

$$c = 1.8\text{ J/g}^\circ\text{C}$$

4. 100.0 g of 4.0°C water is heated until its temperature is 37°C. If the specific heat of water is 4.18 J/g°C, calculate the amount of heat energy needed to cause this rise in temperature.

$$Q = mc\Delta T$$

$$Q = (100)(4.18)(37 - 4)$$

$$Q = 13827\text{ J}$$

Gradeamathhelp Answer Keys

Raffaela Di Napoli

A red circular graphic with a gradient, appearing as a partial circle or a thick arc, located to the right of the author's name.

Gradeamathhelp Answer Keys:

Embracing the Melody of Expression: An Emotional Symphony within **Gradeamathhelp Answer Keys**

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Table of Contents Gradeamathhelp Answer Keys

1. Understanding the eBook Gradeamathhelp Answer Keys
 - The Rise of Digital Reading Gradeamathhelp Answer Keys
 - Advantages of eBooks Over Traditional Books
2. Identifying Gradeamathhelp Answer Keys
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Gradeamathhelp Answer Keys
 - User-Friendly Interface
4. Exploring eBook Recommendations from Gradeamathhelp Answer Keys

- Personalized Recommendations
 - Gradeamathhelp Answer Keys User Reviews and Ratings
 - Gradeamathhelp Answer Keys and Bestseller Lists
5. Accessing Gradeamathhelp Answer Keys Free and Paid eBooks
 - Gradeamathhelp Answer Keys Public Domain eBooks
 - Gradeamathhelp Answer Keys eBook Subscription Services
 - Gradeamathhelp Answer Keys Budget-Friendly Options
 6. Navigating Gradeamathhelp Answer Keys eBook Formats
 - ePub, PDF, MOBI, and More
 - Gradeamathhelp Answer Keys Compatibility with Devices
 - Gradeamathhelp Answer Keys Enhanced eBook Features
 7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Gradeamathhelp Answer Keys
 - Highlighting and Note-Taking Gradeamathhelp Answer Keys
 - Interactive Elements Gradeamathhelp Answer Keys
 8. Staying Engaged with Gradeamathhelp Answer Keys
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Gradeamathhelp Answer Keys
 9. Balancing eBooks and Physical Books Gradeamathhelp Answer Keys
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Gradeamathhelp Answer Keys
 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
 11. Cultivating a Reading Routine Gradeamathhelp Answer Keys
 - Setting Reading Goals Gradeamathhelp Answer Keys
 - Carving Out Dedicated Reading Time
 12. Sourcing Reliable Information of Gradeamathhelp Answer Keys

- Fact-Checking eBook Content of Gradeamathhelp Answer Keys
- Distinguishing Credible Sources

13. Promoting Lifelong Learning

- Utilizing eBooks for Skill Development
- Exploring Educational eBooks

14. Embracing eBook Trends

- Integration of Multimedia Elements
- Interactive and Gamified eBooks

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