

Periodic Table of the Elements

Legend:

- metals
- metalloids
- nonmetals
- halogens
- noble gases
- transition metals
- lanthanide series
- actinide series

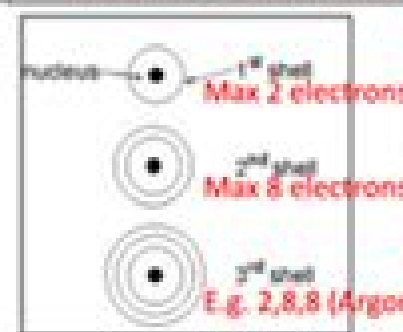
1	2											3	4	5	6	7	8	9	10	11	12	13	14	15	16	17	18																			
H	He											Li	Be	B	C	N	O	F	Ne	Na	Mg	Al	Si	P	S	Cl	Ar																			
K	Ca	Sc	Ti	V	Cr	Mn	Fe	Co	Ni	Cu	Zn	Ga	Ge	As	Se	Br	Kr	Rb	Sr	Y	Zr	Nb	Mo	Tc	Ru	Rh	Pd	Ag	Cd	In	Sn	Sb	Te	I	Xe											
Ba	La	Ce	Pr	Nd	Pm	Sm	Eu	Gd	Tb	Dy	Ho	Er	Tm	Yb	Lu	Hf	Ta	W	Re	Os	Ir	Pt	Au	Hg	Tl	Pb	Bi	Po	At	Fr	Ra	Ac	Th	Pa	U	Np	Pu	Am	Cm	Bk	Cf	Es	Fm	Md	No	Lr

An atom:
Number of Protons = Number of Electrons
Mass number = Protons + Neutrons

C2 Chemistry

Periodic table arranges elements by **ATOMIC NUMBER** (proton number/the small one)

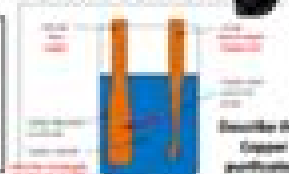
The number of outer shell electrons match the group the element is found in.
E.g. Lithium 2,1 is a group 1 element.



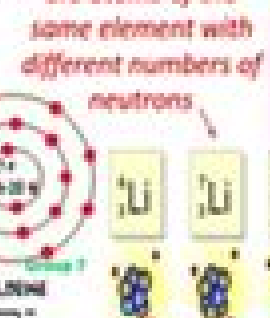
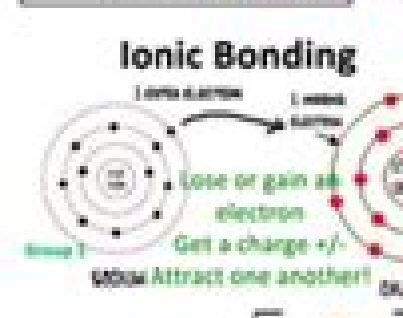
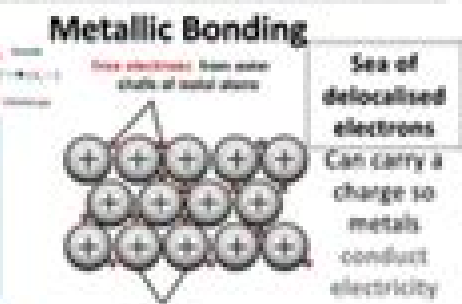
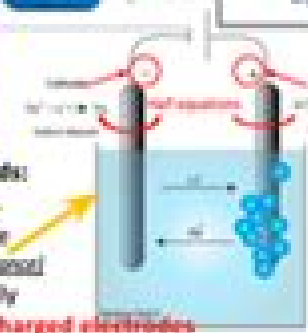
Protons and neutrons both have a mass of 1

ISOTOPES
are atoms of the same element with different numbers of neutrons

OIL RIG
Oxidation is loss (of electrons)
Reduction is gain (of electrons)



Electrolysis:
Splitting up a substance using electricity



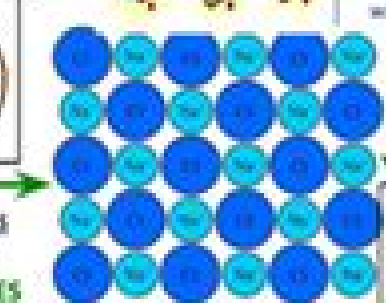
When there are 2 possible products
The less reactive forms at the electrode



Metal ions only arranged in layers
ELECTROSTATIC ATTRACTION
between
+IONS and ELECTRONS

Metal ions slide over one another making them **MALLEABLE**

Metals conduct **HEAT** well



As a solid: High melting and Boiling Point
Conducts electricity when melted or dissolved:
Ions move freely and carry a charge

NANO-TECHNOLOGY
Means **REALLY REALLY REALLY** small!!!!

Gcse Additional Chemistry C2 Revision Booklet

David Baud



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