

Unit 1 • Observations, Models, & Experiments**SCIENTIFIC NOTATION**

Change the following to Scientific Notation (maintain the number of significant figures):

1. 5.280 = 5.280

11. 2,560 = 2.56×10^3

2. 2,000 = 2×10^3

12. .0009 = 9×10^{-4}

3. 15 = 1.5×10^1

13. 8,900,000 = 8.9×10^6

4. 6,589,000 = 6.589×10^6

14. .0920 = 9.20×10^{-2}

5. 70,400,000,000 = 7.04×10^{10}

15. 6,300 = 6.3×10^3

6. .00263 = 2.63×10^{-3}

16. .90 = 9.0×10^{-1}

7. .00589 = 5.89×10^{-3}

17. 250 = 2.5×10^2

8. .006 = 6×10^{-3}

18. .006087 = 6.087×10^{-3}

9. .400 = 4.00×10^{-1}

19. 500,000 = 5×10^5

10. .08060 = 8.060×10^{-2}

20. .0000000105 = 1.05×10^{-8}

Make the following Metric System conversions using "unit analysis" (you may use scientific notation):

1. 100 mg $\times \frac{1 \text{ g}}{1000 \text{ mg}}$ = .1 g

2. 20 cm $\times \frac{1 \text{ m}}{100 \text{ cm}}$ = .2 m

3. 50 L $\times \frac{1 \text{ kL}}{1000 \text{ L}}$ = .05 kL

4. 22 g $\times \frac{100 \text{ cg}}{1 \text{ g}}$ = 2200 cg

5. 825 cm $\times \frac{1 \text{ m}}{100 \text{ cm}} \times \frac{1 \text{ km}}{1000 \text{ m}}$ = 8.25×10^{-3} km

6. 2,350 kg $\times \frac{1000 \text{ g}}{1 \text{ kg}}$ = 2.35×10^6 g

7. 19 mL $\times \frac{1 \text{ L}}{1000 \text{ mL}} \times \frac{100 \text{ cL}}{1 \text{ L}}$ = 1.9 cL

8. 52 km $\times \frac{1000 \text{ m}}{1 \text{ km}}$ = 52000 m

9. 36 m $\times \frac{100 \text{ cm}}{1 \text{ m}}$ = 3600 cm

10. 18 cm $\times \frac{1 \text{ m}}{100 \text{ cm}} \times \frac{1000 \text{ mm}}{1 \text{ m}}$ = 180 mm

11. 6 g $\times \frac{1000 \text{ mg}}{1 \text{ g}}$ = 6000 mg

12. 4,259 mg $\times \frac{1 \text{ g}}{1000 \text{ mg}}$ = 4.259 g

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