

ChemThink: Molecular Shapes

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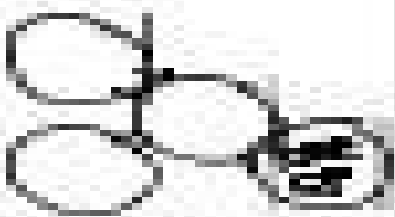

Lesson 8

Electrons

Page

Electron pairs

Understanding that as the number of molecules that have lone pairs on an atom, the electron pairs will repel each other and this will repel and it affect the shape of the molecule.

Number of atoms	Electron pairs	Bond Angles	Shape	Ex. Lewis Dot representation
2		180°	Linear	$H-H$
3		120°	Trigonal Planar	
4		109°	Bent	

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