

Name _____

Molecular Geometries Lab Report Sheet

Formula	Lewis Structure	No of electron domains/Atp	Electron domain geometry	VSEPR Molecular geometry	Hybridization of central atom	Polar ?
CO ₂	$\begin{array}{c} \text{Valence } 24 \\ \text{:O}=\text{C}=\text{O:} \\ \text{2 charge clouds} \end{array}$	0	Linear 180°	Linear 	SP	NP
BF ₃	$\begin{array}{c} \text{Valence } 32 \\ \text{:F}=\text{B}-\text{F:} \\ \text{:F:} \end{array}$	0	Trigonal Planar 120°	Trigonal Planar 	SP ₂	NP
SnCl ₂	$\begin{array}{c} \text{Valence } 24 \\ \text{:Cl}=\text{Sn}-\text{Cl:} \\ \text{:} \end{array}$	3	Bent 120°	Bent 	SP ₂	P
BF ₄ ⁻	$\begin{array}{c} \text{Valence } 32 \\ \left[\begin{array}{c} \text{:F:} \\ \text{:F}-\text{B}-\text{F:} \\ \text{:F:} \end{array} \right]^- \end{array}$	0	Tetrahedral 109.5°	Tetrahedral 	sp ³	P
AsH ₃	$\begin{array}{c} \text{H} \\ \\ \text{H}-\text{As}-\text{H} \\ \\ \text{H} \end{array}$	4	Trigonal Pyramidal 109.5°	Trigonal Pyramidal 	sp ³	P
H ₂ S	$\text{H}-\ddot{\text{S}}-\text{H}$	2	Bent 109.5°	Bent 	sp ³	NP
NH ₂ ⁻	$\left[\text{H}-\ddot{\text{N}}-\text{H} \right]^-$	2	Bent 109.5°	Bent 	sp ³	P
POF ₃	$\begin{array}{c} \text{Valence } 40 \\ \text{:F} \quad \text{:F:} \\ \diagdown \quad / \\ \text{:F}-\text{P}-\text{Cl:} \\ \\ \text{:F:} \end{array}$	0	Trigonal Bipyre 120° + 90°	Trigonal Bipyre 	sp ³ d	P

Molecular Geometry And Polarity Lab Answers

**Kenneth W. Whitten, Carl B.
Bishop, Kenneth D. Gailey**



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