

**CHAPTER 11 REVIEW***Molecular Composition of Gases***MIXED REVIEW****SHORT ANSWER** Answer the following questions in the space provided.

1. \_\_\_\_\_ The average speed of a gas molecule is most directly related to the \_\_\_\_\_.  
(a) polarity of the molecule  
(b) pressure of the gas  
(c) temperature of the gas  
(d) number of moles in the sample
2. \_\_\_\_\_ Which of the following statements best represents Graham's law for two gases at the same temperature?  
(a) The molar masses and the velocities are inversely proportional.  
(b) The molar masses and the velocities are directly proportional.  
(c) The molar masses and the square roots of the velocities are directly proportional.  
(d) The molar masses and the squares of the velocities are inversely proportional.
3. Water is synthesized from its elements according to the following equation:  
$$2\text{H}_2(g) + \text{O}_2(g) \rightarrow 2\text{H}_2\text{O}(l)$$
  
\_\_\_\_\_ a. Which is greater, the volume of  $\text{H}_2$  consumed or the volume of  $\text{O}_2$  consumed?  
\_\_\_\_\_ b. Which is greater, the mass of  $\text{H}_2$  consumed or the mass of  $\text{O}_2$  consumed?
4. Consider a 2.24 L sample of an ideal gas at STP.  
\_\_\_\_\_ a. How many moles are present in the sample, or is more data needed?  
\_\_\_\_\_ b. How many molecules are present in the sample, or is more data needed?  
\_\_\_\_\_ c. What is the mass of the sample, or is more data needed?

**PROBLEMS** Write the answer on the line to the left. Show all your work in the space provided.

5. \_\_\_\_\_ Calculate the volume of 3.5 mol of an ideal gas at  $24^\circ\text{C}$  and 760 torr.

# Molecular Composition Of Gases 11 3 Answers

**Patrick Vollmar**



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