

Name \_\_\_\_\_ answer key \_\_\_\_\_ Date \_\_\_\_\_

# Stoichiometry: Mass – Mole Worksheet

Based on the following chemical equation, please answer the following questions!



Calculate the amount of moles of NaCl needed to produce 75 g of FeCl<sub>3</sub>.

9 FeCl <sub>3</sub>	1 mole FeCl <sub>3</sub>	mole NaCl
9 FeCl <sub>3</sub>	9 FeCl <sub>3</sub>	mole FeCl

Calculate the amount of Fe<sub>2</sub> (SO<sub>4</sub>)<sub>3</sub> in grams produce 215 g Na<sub>2</sub>SO<sub>4</sub>.

9 Na <sub>2</sub> SO <sub>4</sub>	1 mole Na <sub>2</sub> SO <sub>4</sub>	mole Fe <sub>2</sub> (SO <sub>4</sub> ) <sub>3</sub>	9 Fe <sub>2</sub> (SO <sub>4</sub> ) <sub>3</sub>
9 Na <sub>2</sub> SO <sub>4</sub>	9 Na <sub>2</sub> SO <sub>4</sub>	mole Na <sub>2</sub> SO <sub>4</sub>	1 mole Fe <sub>2</sub> (SO <sub>4</sub> ) <sub>3</sub>

Calculate the amount in grams of FeCl<sub>3</sub> produced with 95g of Fe (SO<sub>4</sub>)<sub>3</sub>.

9 Fe <sub>2</sub> (SO <sub>4</sub> ) <sub>3</sub>	1 mole Fe <sub>2</sub> (SO <sub>4</sub> ) <sub>3</sub>	mole FeCl <sub>3</sub>	9 FeCl <sub>3</sub>
9 Fe <sub>2</sub> (SO <sub>4</sub> ) <sub>3</sub>	9 Fe <sub>2</sub> (SO <sub>4</sub> ) <sub>3</sub>	mole Fe(SO <sub>4</sub> ) <sub>3</sub>	1 mole FeCl <sub>3</sub>

Calculate the amount of moles of Fe<sub>2</sub>(SO<sub>4</sub>)<sub>3</sub> needed to react with 123.5 g of NaCl

9 NaCl	1 mole NaCl	mole Fe(SO <sub>4</sub> ) <sub>3</sub>
9 NaCl	9 NaCl	mole NaCl

Based on the following chemical equation, please answer the following questions!



9 Ca(HCO <sub>3</sub> ) <sub>2</sub>	1 mole Ca(HCO <sub>3</sub> ) <sub>2</sub>	mole CaCO <sub>3</sub>
9 Ca(HCO <sub>3</sub> ) <sub>2</sub>	9 Ca(HCO <sub>3</sub> ) <sub>2</sub>	mole Ca(HCO <sub>3</sub> ) <sub>2</sub>

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