

Molarity Practice

Calculate for the unknown variables. $M = n/v$

1. Explain what molarity/concentration means?
2. What is the molarity of a solution that contains 5.50 moles of calcium chloride in 2.5 liters of solution?
3. What is the molarity of a solution that contains 3.39×10^{-2} moles hydrobromic acid in 1 liter of solution?
4. A beaker contains 250g $C_6H_{12}O_6$ (glucose) in 1,500 mL of solution. What is the molarity?
5. A flask contains 325g grams iron II oxide and has a molarity of 2M. What is the volume of the solution?
6. Calculate the moles if a flask contains 3.25M potassium nitrate in 750 mL of solution.
7. A cup contains 85.31 grams ammonium acetate in 3.27 liters of solution - calculate the molarity.
8. What is the molarity of a solution that contains 10.5 grams magnesium oxide in 250 mL of solution?
9. What mass of potassium phosphate would you use to make 2.5 liter of 1.25 M solution?

Molarity Practice Unit Lesson 2

Anthony A. Gianelly

A red circular graphic with a gradient, appearing as a partial circle or a thick curved line, located to the right of the author's name.

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A Problem-Solving Approach to Aquatic Chemistry James N. Jensen, 2023-01-05 A Problem Solving Approach to Aquatic Chemistry Enables civil and environmental engineers to understand the theory and application of aquatic equilibrium chemistry The second edition of A Problem Solving Approach to Aquatic Chemistry provides a detailed introduction to aquatic equilibrium chemistry calculation methods for systems at equilibrium applications of aquatic chemistry and chemical kinetics The text directly addresses two required ABET program outcomes in environmental engineering chemistry including stoichiometry equilibrium and kinetics and material and energy balances fate and transport of substances in and between air water and soil phases The book is very student centered with each chapter beginning with an introduction and ending with a summary that reviews the chapter s main points To aid in reader comprehension important terms are defined in context and key ideas are summarized Many thought provoking discussion questions worked examples and end of chapter problems are also included Each part of the text begins with a case study a portion of which is addressed in each subsequent chapter illustrating the principles of that chapter In addition each chapter has an Historical Note exploring connections with the people and cultures connected to topics in the text A Problem Solving Approach to Aquatic Chemistry includes Fundamental concepts such as concentration units thermodynamic basis of equilibrium and manipulating equilibria Solutions of chemical equilibrium problems including setting up the problems and algebraic graphical and computer solution techniques Acid base equilibria including the concepts of acids and bases titrations and alkalinity and acidity Complexation including metals ligands equilibrium calculations with complexes and applications of complexation chemistry Oxidation reduction equilibria including equilibrium calculations graphical approaches and applications Gas liquid and solid liquid equilibrium with expanded coverage of the effects of global climate change Other topics including chemical kinetics of aquatic systems surface chemistry and integrative case studies For advanced senior undergraduates and first year graduate students in environmental engineering courses A Problem Solving Approach to Aquatic Chemistry serves as an invaluable learning resource on the topic with a variety of helpful learning elements included throughout to ensure information retention and the ability to apply covered concepts in practical settings

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standards **Loose-Leaf Version for Chemical Principles** Peter Atkins, Loretta Jones, Leroy Laverman, 2016-01-07
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because their f electron shells are gradually filled as the atomic number increases 4f for lanthanides and 5f for actinides 1 these elements are known as f block elements figure 1

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