

- (i) It depends on the concentration of a non-electrolyte solute in a solution and does not depend on the nature of the solute molecules.
(ii) It depends on a number of particles of electrolyte solute in a solution and does not depend on the nature of the solute particles.
(iii) It depends on the concentration of a non-electrolyte solute in solution as well as on the nature of the solute molecules.
(iv) It depends on the concentration of an electrolyte or nonelectrolyte solute in solution as well as on the nature of solute molecules.

Solution:

Option (i) and (ii) are the answers.

30. Van't Hoff factor i is given by the expression _____.

- (i) $i = \text{Normal molar mass} / \text{Abnormal molar mass}$
(ii) $i = \text{Abnormal molar mass} / \text{Normal molar mass}$
(iii) $i = \text{Observed colligative property} / \text{calculated colligative property}$
(iv) $i = \text{Calculated colligative property} / \text{Observed colligative property}$

Solution:

Option (i) and (iii) are the answers.

31. Isotonic solutions must have the same _____.

- (i) solute
(ii) density
(iii) elevation in boiling point
(iv) depression in freezing point

Solution:

Option (iii) and (iv) are the answers.

32. Which of the following binary mixtures will have the same composition in the liquid and vapour phase?

- (i) Benzene - Toluene
(ii) Water-Nitric acid
(iii) Water-Ethanol
(iv) n -Hexane - n -Heptane

Solution:

Option (ii) and (iii) are the answers.

33. In isotonic solutions _____.

- (i) solute and solvent both are same.
(ii) osmotic pressure is the same.
(iii) solute and solvent may or may not be same.
(iv) a solute is always the same solvent may be different.

Solution:

Option (ii) and (iii) are the answers.

34. For a binary ideal liquid solution, the variation in total vapour pressure versus the composition of the solution is given by which of the curves?

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Annelies Wilder-Smith



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