

1. Consider the following reaction:



An equilibrium mixture of this reaction at a certain temp was found to have  $[\text{CO}] = 0.115\text{M}$ ,  $[\text{H}_2] = 0.110\text{M}$ , and  $[\text{CH}_3\text{OH}] = 0.190\text{M}$ . What is the value of the equilibrium constant ( $K_c$ ) at this temperature?

$$K_c = \frac{[\text{CH}_3\text{OH}]}{[\text{CO}][\text{H}_2]^2} = \frac{(0.19)}{(0.115)(0.11)^2} = 136.5$$

2. Consider the following reaction:



An equilibrium mixture of this reaction at a certain temperature was found to have  $[\text{NH}_3] = 0.274\text{M}$  and  $[\text{H}_2\text{S}] = 0.355\text{M}$ . What is the value of the equilibrium constant ( $K_c$ ) at this temperature?

$$K_c = [\text{NH}_3][\text{H}_2\text{S}] = (0.274)(0.355) = 0.0973$$

3. Consider the following reaction:



The reaction between CO and  $\text{H}_2$  is carried out at a specific temperature with initial concentrations of  $\text{CO} = 0.32\text{M}$  and  $\text{H}_2 = 0.45\text{M}$ . At equilibrium, the concentration of  $\text{CH}_3\text{OH}$  is  $0.14\text{M}$ . Find the equilibrium constant at this temperature.

	CO	$2\text{H}_2$	$\text{CH}_3\text{OH}$
I	0.32	0.45	
C	-x	-2x	+x
E	0.18	0.17	0.14

$$K_c = \frac{(0.14)}{(0.18)(0.17)^2} = 26.9$$

4. Consider the following reaction:



The reaction of  $\text{CH}_4$  is carried out at some temperature with an initial concentration of  $[\text{CH}_4] = 0.090\text{M}$ . At equilibrium, the concentration of  $\text{H}_2$  is  $0.012\text{M}$ . Find the equilibrium constant at this temperature.

	$2\text{CH}_4$	$\text{C}_2\text{H}_2$	$3\text{H}_2$
I	0.09		
C	-2x	+x	+3x
E	0.082	0.004	0.012M

$$K_c = \frac{(0.004)(0.012)^3}{(0.082)^2} = 1.0 \times 10^{-6}$$

5. Consider the following reaction:



A reaction mixture contains  $[\text{NO}_2] = 0.0248\text{M}$  and  $[\text{N}_2\text{O}_4] = 0.0337\text{M}$ . Calculate  $Q_c$  at the given concentrations.

$$Q_c = \frac{(0.0248)^2}{(0.0337)} = 0.0183 \quad Q_c > K_c, \text{ rxn goes left}$$

6. The following reaction was carried out in a 2.75L reaction vessel at  $1100\text{K}$ :



# Mastering Chemistry Answers Chapter 19

**CHANG**



## **Mastering Chemistry Answers Chapter 19:**

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**Merrill Chemistry** Robert C. Smoot, Smoot, Richard G. Smith, Jack Price, 1998 *Your Coach (in a Book)* Robert Hargrove, Michel Renaud, 2004-06-28 Coaching has proven to be one of the most power and effective ways for leaders to develop and improve their performance Yet working one on one with a coach is not always possible If you want the experience of masterful coaching Your Coach In a Book provides a time efficient and affordable solution Based on the authors highly successful Masterful Coaching approach Your Coach In a Book is designed to help you master your trickiest leadership business and career challenges Throughout master level coaches Robert Hargrove and Michel Renaud engage the reader in a coaching conversation about your most important goals pivotal decisions bothersome issues and dilemmas Your Coach In A Book simulates the experience of working with a personal coach drawing on thousands of hours of coaching conversations It gives you the insights you need to set aspirational goals master the corporate chessboard and create new openings for action where you are stuck or ineffective

*End-User Considerations in Educational Technology Design* Roscoe, Rod D., Craig, Scotty D., Douglas, Ian, 2017-06-16 Emerging technologies have enhanced the learning capabilities and opportunities in modern school systems To continue the effective development of such innovations the intended users must be taken into account End User Considerations in Educational Technology Design is a pivotal reference source for the latest scholarly material on usability testing techniques and user centered design methodologies in the development of technological tools for learning environments Highlighting a range of pertinent topics such as multimedia learning human computer interaction and

online learning this book is ideally designed for academics researchers school administrators professionals and practitioners interested in the design of optimized educational technologies      Organic Chemistry as a Second Language David R. Klein,2024-04-02 Organic chemistry can be a challenging subject Most students view organic chemistry as a subject requiring hours upon hours of memorization Author David Klein s Second Language books prove this is not true organic chemistry is one continuous story that makes sense if you pay attention Offering a unique skill building approach these market leading books teach students how to ask the right questions to solve problems study more efficiently to avoid wasting time and learn to speak the language of organic chemistry Covering the initial half of the course Organic Chemistry as a Second Language First Semester Topics reviews critical principles and explains their relevance to the rest of the course Each section provides hands on exercises and step by step explanations to help students fully comprehend classroom lectures and textbook content Now in the 6th edition there are approximately 30 new end of chapter exercises in each chapter These new exercises vary in difficulty starting with exercises that focus on just one skill or concept called Practice Problems and continuing with exercises that focus on more than one skill or concept called Integrated Problems and concluding with advanced exercises called Challenge Problems There are also author created detailed solutions for all new exercises and these detailed solutions appear in the back of the book      **Mastering Organic Chemistry** Gerald W. Gibson,1979

**Chemistry: The Central Science** Theodore L. Brown,H. Eugene LeMay Jr.,Bruce E. Bursten,Catherine Murphy,Patrick Woodward,Steven Langford,Dalius Sagatys,Adrian George,2013-10-04 If you think you know the Brown LeMay Bursten Chemistry text think again In response to market request we have created the third Australian edition of the US bestseller Chemistry The Central Science An extensive revision has taken this text to new heights Triple checked for scientific accuracy and consistency this edition is a more seamless and cohesive product yet retains the clarity innovative pedagogy functional problem solving and visuals of the previous version All artwork and images are now consistent in quality across the entire text And with a more traditional and logical organisation of the Organic Chemistry content this comprehensive text is the source of all the information and practice problems students are likely to need for conceptual understanding development of problem solving skills reference and test preparation      *Chemistry II For Dummies* John T. Moore,2012-07-03 The tools you need to ace your Chemisty II course College success for virtually all science computing engineering and premedical majors depends in part on passing chemistry The skills learned in chemistry courses are applicable to a number of fields and chemistry courses are essential to students who are studying to become nurses doctors pharmacists clinical technicians engineers and many more among the fastest growing professions But if you re like a lot of students who are confused by chemistry it can seem like a daunting task to tackle the subject That s where Chemistry II For Dummies can help Here you ll get plain English easy to understand explanations of everything you ll encounter in your Chemistry II class Whether chemistry is your chosen area of study a degree requirement or an elective you ll get the skills and confidence to score high

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Physics John Wilkinson,2007 Revision guide for VCE students covering complete course for Units 3 4 as well as 3 options in each Unit It is an effective study program a month before each exam Tells you what to study each week how much time to spend on each section Includes two trial exams one for mid year Exam 1 Unit 3 one for end of year Exam 2 Unit 4

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