

## PHYSICAL SCIENCE: Molecular Theory

### Learning Objectives

- 1) Students should be able to explain the differences between the three states of matter and the transitions between them and understand the molecular theory of matter.
- 2) Students should be able to explain the differences between the three states of matter and the transitions between them.
- 3) Students should be able to explain the differences between the three states of matter and the transitions between them.

### Module 1: Representation of Matter in Different Phases

The following figure shows the representation of matter in three different phases: Solid, Liquid, and Gas.



Figure 1: Representation of Matter in Three Different Phases

Students should be able to explain the differences between the three states of matter and the transitions between them and understand the molecular theory of matter.

### Key Concepts


1. Matter can be in three different states: Solid, Liquid, and Gas. The particles in matter are in constant motion and the state of matter depends on the temperature and pressure.
2. The particles in matter are in constant motion and the state of matter depends on the temperature and pressure.
3. The particles in matter are in constant motion and the state of matter depends on the temperature and pressure.

Table 1: Representation of Matter in Three Different Phases

	Solid	Liquid	Gas
<b>Shape and Volume</b>	Definite shape and volume. The particles are in a regular, packed arrangement.	Definite volume but no definite shape. The particles are in a disordered, close-packed arrangement.	No definite shape or volume. The particles are in a highly disordered, spread-out arrangement.
<b>Particle Arrangement</b>	Particles are in a regular, packed arrangement. They are in constant motion but their relative positions are fixed.	Particles are in a disordered, close-packed arrangement. They are in constant motion and their relative positions are not fixed.	Particles are in a highly disordered, spread-out arrangement. They are in constant motion and their relative positions are not fixed.
<b>Intermolecular Forces</b>	Strong intermolecular forces. The particles are held together tightly.	Weak intermolecular forces. The particles are held together loosely.	Very weak intermolecular forces. The particles are held together very loosely.

# Kinetic Molecular Theory Pogil Instructors Guide

**Ben B. Chastain, Richard Earl  
Dickerson, Harry B. Gray, Gilbert Pierce  
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