

Example:

25. Three grams of carbon and 15 g of oxygen are placed in a closed vessel and heated to 1000°C. What mass of CO_2 is formed?

$$25 \text{ g } \text{C} \times \frac{1 \text{ mol C}}{12 \text{ g C}} \times \frac{8 \text{ mol CO}_2}{1 \text{ mol C}} \times \frac{44 \text{ g CO}_2}{1 \text{ mol CO}_2} = 73.3 \text{ g CO}_2$$

26. When 10 g of aluminum reacts with oxygen, 20 g of Al_2O_3 is formed. How much oxygen is required to form 20 g of Al_2O_3 ?



$$10 \text{ g Al} \times \frac{1 \text{ mol Al}}{27 \text{ g Al}} \times \frac{2 \text{ mol Al}_2\text{O}_3}{4 \text{ mol Al}} \times \frac{102 \text{ g Al}_2\text{O}_3}{1 \text{ mol Al}_2\text{O}_3} = 37.0 \text{ g Al}_2\text{O}_3 \text{ (theoretical yield)}$$

$$\% \text{ yield} = \frac{20 \text{ g Al}_2\text{O}_3}{37.0 \text{ g Al}_2\text{O}_3} \times 100 = 54.1\%$$

27. In the reaction of problem 26, what is the limiting reagent when 10 g of Al and 10 g of O_2 are used?

$$10 \text{ g Al} \times \frac{1 \text{ mol Al}}{27 \text{ g Al}} \times \frac{2 \text{ mol Al}_2\text{O}_3}{4 \text{ mol Al}} \times \frac{102 \text{ g Al}_2\text{O}_3}{1 \text{ mol Al}_2\text{O}_3} = 37.0 \text{ g Al}_2\text{O}_3$$

$$10 \text{ g O}_2 \times \frac{1 \text{ mol O}_2}{32 \text{ g O}_2} \times \frac{2 \text{ mol Al}_2\text{O}_3}{3 \text{ mol O}_2} \times \frac{102 \text{ g Al}_2\text{O}_3}{1 \text{ mol Al}_2\text{O}_3} = 20.6 \text{ g Al}_2\text{O}_3 \text{ (actual product)}$$

Limiting Reagent = O_2

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Hrw

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