Chapter 6 Practice Test: Chemical Bonding

In the space provided, write the letter of the term or phrase that best completes each statement or best answers each question.

- 1. The charge on an ion is
 - a. always positive.
 - b. always negative.
 - c. either positive or negative.
 - d. zero.
- According to the octet rule, a calcium atom has a tendency to
 - a. lose one electron.
 - b. lose two electrons.
 - c. gain one electron.
 - d. gain two electrons.
- 3. If a compound forms by ionic bonding, which is not true?
 - a. A positively charged atom or group of atoms attracts a negatively charged atom or group of atoms.
 - The net charge of the compound is zero.
 - c. The compound contains just two atoms, each of opposite charge.
 - Several ions group together in a tightly packed structure.
- _____ 4. The only property listed that is not characteristic of ionic compounds is
 - a. high melting point.
 - b. hardness.
 - c. lack of crystal structure.
 - d. good conductor of electricity.
- 5. Which formula listed below represents a polyatomic ion?
 - a. NH₄*
 - b. H₂SO₄
 - c. NaCl.
 - d. H₂O
- 6. The melting points of ionic compounds are higher than the melting points of molecular compounds because
 - a. ionic substances tend to vaporize at room temperature.
 - b. ionic substances are brittle.
 - attractive forces between ions are greater than the attractive forces between molecules.
 - d. the numbers of positive and negative charges are equal in an ionic compound.

Holt Chemistry Chapter 6 Test

David W. Oxtoby, H. Pat Gillis, Laurie J. Butler

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and molecular orbital theory are introduced after the student has developed an understanding of fundamental trends in chemical properties and reactions across the periodic table which allows MO theory to be more broadly applied in subsequent chapters Use of this text is expected to increase student enrollment and build students appreciation of the central role of inorganic chemistry in any allied field Key Features Over 900 end of chapter exercises half answered in the back of the book Over 180 worked examples Optional experiments demos Clearly cited connections to other areas in chemistry and chemical sciences Chapter opening biographical vignettes of noted scientists in Inorganic Chemistry Optional General Chemistry review sections Originally rendered two color illustrations throughout Essentials of Modern Chemistry Charles Elwood Dull, 1918 Modern Methods of Drug Discovery Alexander Hillisch, Rolf Hilgenfeld, 2012-11-28 Research in the pharmaceutical industry today is in many respects quite different from what it used to be only fifteen years ago There have been dramatic changes in approaches for identifying new chemical entities with a desired biological activity While chemical modification of existing leads was the most important approach in the 1970s and 1980s high throughput screening and structure based design are now major players among a multitude of methods used in drug discov ery Quite often companies favor one of these relatively new approaches over the other e g screening over rational design or vice versa but we believe that an intelligent and concerted use of several or all methods currently available to drug discovery will be more successful in the medium term What has changed most significantly in the past few years is the time available for identifying new chemical entities Because of the high costs of drug discovery projects pressure for maximum success in the shortest possible time is higher than ever In addition the multidisciplinary character of the field is much more pronounced today than it used to be As a consequence researchers and project managers in the pharmaceutical industry should have a solid knowledge of the more important methods available to drug discovery because it is the rapidly and intelligently combined use of these which will determine the success or failure of preclinical projects Laboratory Experiments to Accompany Lab Experiments Modern Chemistry Harold Clark Metcalfe, 1982 "Modern Chemistry," Fredus Nelson Peters, 1902

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