Name

Steichiometry

- How do you determine the amount of product formed in a reaction?
 Use stoichiometry based on the limiting reactant.
- 22. What is the limiting reactant in $2SO_7*O_7*2SO_3$ if you start with 3 motes of SO_2 and 2 motes of O_2 ?

5022

23. Why is it important to balance a chemical equation for stoichiometry calculations?

To ensure the correct mole ratios are used.

24. How can you increase the percent yield of a reaction?

Optimize reaction conditions and minimize losses.

25. What information is needed to calculate theoretical yield?

Balanced chemical equation and amounts of reactants.

26. If the percent yield is 50% and the actual yield is 5 grams, what is the theoretical yield?

10 grams

- What is the first step in stoichiometric calculations?
 Astance the chemical equation.
- 28. In C_2H_3 +5 O_2 +3 CO_2 +4 H_2O , what is the mote ratio of O_2 to CO_2 ?

 5:3
- 29. If you have 10 grams of NaCl and the reaction has a percent yield of 75%, how much product is obtained?

7.5 grams

30. How do you identify the limiting reactant in a multi-step reaction?

Determine the limiting reactant for each step and use it for calculations in subsequent steps.



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RJ Shavelson

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