

## Practice Test #15: Chemical Equilibrium

1. Which one of the following will change the value of an equilibrium constant?

- A) Changing the temperature
- B) Adding other substances that do not react with any of the species involved in the equilibrium
- C) Varying the initial concentrations of reactants
- D) Varying the initial concentrations of products

2. An aqueous equilibrium mixture of  $\text{CoCl}_4^{2-}$ ,  $\text{CoBr}_4^{2-}$ ,  $\text{Cl}^-$ , and  $\text{Br}^-$  is present in a flask at  $25^\circ\text{C}$ . Which action below, will change the value of the equilibrium constant from that which currently describes the concentration relationships of the four species above:

- A) add more  $\text{Cl}^-$  to the solution
- B) add more  $\text{Br}^-$  to the solution
- C) add more  $\text{CoBr}_4^{2-}$  to the solution
- D) add more  $\text{CoCl}_4^{2-}$  to the solution
- E) put the flask into an  $80^\circ\text{C}$  water bath

3. The equilibrium constant for reaction (1) is "K". What is the equilibrium constant for equation (2)?



- A)  $K^2$
- B)  $2K$
- C)  $\frac{1}{2} K$
- D)  $1/K^2$

The second equation is both flipped and multiplied by 2, resulting in the value of K being inverted and squared.

4. The value of K for the following reaction is 0.25.



What is the value of K for the reaction shown below?



- A) 0.50
- B) 0.0625
- C) 0.12
- D) 0.25

The second equation is the same as the first except that it is multiplied by 2.

The new value of K is  $0.25^2 = 0.0625$

# Holt Chemistry Chemical Equilibrium Answer Key Test

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