#### Chapter 2 The Chemistry of Biology

#### Atoms, Bonds, and Molecules

: all tangible materials that occupy space and have mass

Atom: a tiny particle that cannot be subdivided into smaller substances without losing its properties

Properties of atoms:

- Protons (p\*): positively charged, found in the nucleus
- Neutrons (n<sup>0</sup>): no charge, also found in the nucleus
- Electrons (e'): negatively charged, surround the nucleus

#### **Electron Orbitals and Shells**

Atomic number of protons
Atomic mass: number of protons and neutrons

Orbidal

Pathways in which electrons rotate around the nucleus

#### Models of Atomic Structure

in image

#### Major Elements of Life and Their Primary

#### Characteristics 1

Element :

An atom with a characteristic atomic structure and predictable chemical behavior

#### **Bonds and Molecules 1**

Molecule

- A chemical substance that results from the combination of two or more atoms
- O<sub>2</sub>, N<sub>2</sub>

Compound

- A molecule composed of two or more different elements
- H<sub>2</sub>O<sub>2</sub> C<sub>6</sub>H<sub>12</sub>O<sub>6</sub>

#### Bonds and Molecules 2

Chemical bonds:

Result when two or more atoms share, donate (lose), or accept (gain) electrons

#### Covalent Bonds and Polarity 1

Covalent bonds

- Form between atoms that share electrons rather than donating or receiving them
- Single covalent bonds share a pair of electrons (H<sub>2</sub>)
- Double \_\_\_\_\_ covalent bonds share two pairs of electrons and are more rigid

# **Holt Biology Chapter 2c Test**

James J. Heckman, John Eric Humphries, Tim Kautz **Holt Biology Chapter 2c Test:** 

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