



ESSENTIALS

OF FIRE FIGHTING
AND FIRE DEPARTMENT OPERATIONS



Test Bank

Essentials Of Firefighting 5th Edition Test Bank

Ed Bowker Staff, Staff Bowker, Ed



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 0.444 mol of COCl_2 in 0.654 l of solution
 98.0 g of phosphoric acid H_3PO_4 in 1.00 l of solution
 0.2074 g of calcium hydroxide Ca(OH)_2 in 40.00 ml of solution

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 a 0.195 g of cholesterol $\text{C}_{27}\text{H}_{46}\text{O}$ in 0.100 l of serum the average concentration of cholesterol in human serum
 b 4.25 g of NH_3 in 0.500 l of solution the concentration of NH_3 in household ammonia

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 molality = $\frac{\text{moles solute}}{\text{kilograms solvent}}$
 mathematical manipulation of molality is the

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 $\frac{0.825 \text{ g}}{141.9579 \text{ g/mol}} \times 0.0129 \text{ m} = 0.00060581158 \text{ mol}$
 $\frac{0.00060581158 \text{ mol}}{0.4500 \text{ kg}} = 0.001346247111 \text{ mol/kg}$
 3 mole fraction Na_2HPO_4
 $\frac{0.825 \text{ g}}{141.9579 \text{ g/mol}} = 0.00581158 \text{ mol}$
 $\frac{0.00581158 \text{ mol}}{0.4500 \text{ g}} = 0.01291462222 \text{ mol/g}$
 $\frac{0.01291462222 \text{ mol/g}}{1000} = 0.00001291462222 \text{ mol/kg}$

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$$\begin{array}{l} \text{mass of a component} \\ \text{frac mass of component in solution total mass of} \\ \text{solution} \times 100 \end{array}$$

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