



Chapter 4 & 5: Modern Model of the Atom and Bonding

Supplemental Instruction
Iowa State University

Leader:	Katie
Course:	Chem 163
Instructor:	Burnett/Kingston
Date:	September 28, 2009

Review

- What is the noble gas configuration of the following atoms and ions?
 - Ar $[\text{Ne}] 3s^2 3p^4$ or $[\text{Ar}]$
 - Na^+ $[\text{Ne}] 2s^2 2p^6$ or $[\text{Ne}]$
 - C^{2-} $[\text{He}] 2s^2 2p^4$
 - O^{2-} $[\text{He}] 2s^2 2p^6$ or $[\text{Ne}]$
 - Ca^{2+} $[\text{Ar}] 3s^2 3p^6$ or $[\text{Ar}]$

New Material

- Define the following terms and answer associated questions.
 - Heisenberg uncertainty principle: Principle that states that the position and momentum of a subatomic particle cannot be known simultaneously.
 - orbital: 3D volume in space, cloud like distribution of electrons.
 - What is its maximum capacity? $2e^-$
 - covalent bond: Chemical bond where the electrons are shared.
 - What does it represent? The covalent bond represents an energy minimum.
 - Is energy absorbed or released when a covalent bond is formed? Energy is released.
 - bond length:
 - What causes bond length? The attraction of electrons to the nuclei and repulsion of 2 nuclei.
 - What units is bond length measured in? $\text{\AA}$ = Angstrom = 10^{-10}m = 1m
- How many orbitals are there in any s sublevel? Any p sublevel? Any d sublevel?

$s = 1 \text{ orbital} = 2e^-$
 $p = 3 \text{ orbitals} = 6e^-$
 $d = 5 \text{ orbitals} = 10e^-$
- Draw the orbitals for the $2p^5$. How many electrons are there in each orbital?



- Write the following atoms in Lewis Dot notation.
 - N $\text{N}^\bullet$
 - H $\text{H}^\bullet$
 - Xe $\text{Xe}^\bullet$
 - Fe $\text{Fe}^\bullet$

Supplemental Instruction

Modern Chemistry Chapter 12 3 Review Answers

Y Pai



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Table of Contents Modern Chemistry Chapter 12 3 Review Answers

1. Understanding the eBook Modern Chemistry Chapter 12 3 Review Answers
 - The Rise of Digital Reading Modern Chemistry Chapter 12 3 Review Answers
 - Advantages of eBooks Over Traditional Books
2. Identifying Modern Chemistry Chapter 12 3 Review Answers
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Modern Chemistry Chapter 12 3 Review Answers
 - User-Friendly Interface
4. Exploring eBook Recommendations from Modern Chemistry Chapter 12 3 Review Answers

- Personalized Recommendations
 - Modern Chemistry Chapter 12 3 Review Answers User Reviews and Ratings
 - Modern Chemistry Chapter 12 3 Review Answers and Bestseller Lists
5. Accessing Modern Chemistry Chapter 12 3 Review Answers Free and Paid eBooks
- Modern Chemistry Chapter 12 3 Review Answers Public Domain eBooks
 - Modern Chemistry Chapter 12 3 Review Answers eBook Subscription Services
 - Modern Chemistry Chapter 12 3 Review Answers Budget-Friendly Options
6. Navigating Modern Chemistry Chapter 12 3 Review Answers eBook Formats
- ePub, PDF, MOBI, and More
 - Modern Chemistry Chapter 12 3 Review Answers Compatibility with Devices
 - Modern Chemistry Chapter 12 3 Review Answers Enhanced eBook Features
7. Enhancing Your Reading Experience
- Adjustable Fonts and Text Sizes of Modern Chemistry Chapter 12 3 Review Answers
 - Highlighting and Note-Taking Modern Chemistry Chapter 12 3 Review Answers
 - Interactive Elements Modern Chemistry Chapter 12 3 Review Answers
8. Staying Engaged with Modern Chemistry Chapter 12 3 Review Answers
- Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Modern Chemistry Chapter 12 3 Review Answers
9. Balancing eBooks and Physical Books Modern Chemistry Chapter 12 3 Review Answers
- Benefits of a Digital Library
 - Creating a Diverse Reading Collection Modern Chemistry Chapter 12 3 Review Answers
10. Overcoming Reading Challenges
- Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
11. Cultivating a Reading Routine Modern Chemistry Chapter 12 3 Review Answers
- Setting Reading Goals Modern Chemistry Chapter 12 3 Review Answers
 - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Modern Chemistry Chapter 12 3 Review Answers

- Fact-Checking eBook Content of Modern Chemistry Chapter 12 3 Review Answers
- Distinguishing Credible Sources

13. Promoting Lifelong Learning

- Utilizing eBooks for Skill Development
- Exploring Educational eBooks

14. Embracing eBook Trends

- Integration of Multimedia Elements
- Interactive and Gamified eBooks

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