

## The Ideal Gas Law

### Introduction

In this lab report, I will be using an experiment to study the ideal gas law. I will use Boyle's Law, Charles's Law, and Avogadro's Law. I will also use the ideal gas law to find the molar mass of a gas.

### The Ideal Gas Law

$$PV = nRT$$

The ideal gas law is a mathematical equation that relates the pressure, volume, and temperature of a gas to the number of particles. The ideal gas law is a good approximation of the behavior of a gas at low pressures and high temperatures. The ideal gas law is not applicable to gases at high pressures and low temperatures.

The ideal gas law is a good approximation of the behavior of a gas at low pressures and high temperatures. The ideal gas law is not applicable to gases at high pressures and low temperatures.



# Ideal Gas Laws Lab Escience

**RC Schank**

A red circular graphic with a gradient, appearing as a semi-circle or a partial circle, located to the right of the RC Schank text.

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