(2)

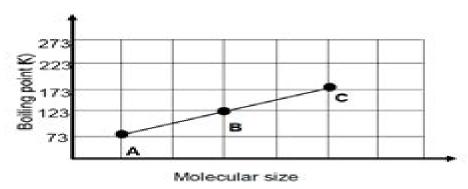
(1)

(3) [13]

QUESTION 2

2.1 The graph of molecular size versus the boiling point is given below.

The **letters A**, **B** and **C** represent the compounds CH₄, C₄H₆ and C₃H₈ respectively.



- 2.1.1 Define the term boiling point.
- 2.1.2. Describe the trend in the boiling points of the compounds as shown by the graph. (2)
- 2.1.3. Explain the answer to QUESTION 2.1.2 by referring to MOLECULAR SIZE, TYPE and STRENGTH of INTERMOLECULAR FORCES. (3)
- 2.1.4 Which ONE of the compounds (A, B or C) has the HIGHEST vapour pressure? Explain the answer by referring to the data on the graph.
 (2)
- 2.2 Consider the two molecules in the table below.

| NAME OFSUBSTANCE | FORMULA | MOLACULAR MASS (g) | BOILING POINT |
|---------------------|-----------------|-----------------------|---------------|
| Ammonia | NH ₃ | 17 | - 33 |
| Phosphine | PH ₃ | × | - 87,4 |

- 2.2.1 Determine the molecular mass (x) of phosphine
- 2.2.2 Explain the difference in the boiling points by referring to the TYPE and STRENGTH of INTERMOLECULAR FORCES.

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Rachel Sandford

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Table of Contents Grade 11 Past Papers Physical Science

- 1. Understanding the eBook Grade 11 Past Papers Physical Science
 - The Rise of Digital Reading Grade 11 Past Papers Physical Science
 - Advantages of eBooks Over Traditional Books
- 2. Identifying Grade 11 Past Papers Physical Science
 - Exploring Different Genres
 - $\circ\,$ Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
- 3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Grade 11 Past Papers Physical Science
 - User-Friendly Interface
- 4. Exploring eBook Recommendations from Grade 11 Past Papers Physical Science

- Personalized Recommendations
- Grade 11 Past Papers Physical Science User Reviews and Ratings
- Grade 11 Past Papers Physical Science and Bestseller Lists
- 5. Accessing Grade 11 Past Papers Physical Science Free and Paid eBooks
 - Grade 11 Past Papers Physical Science Public Domain eBooks
 - Grade 11 Past Papers Physical Science eBook Subscription Services
 - Grade 11 Past Papers Physical Science Budget-Friendly Options
- 6. Navigating Grade 11 Past Papers Physical Science eBook Formats
 - o ePub, PDF, MOBI, and More
 - Grade 11 Past Papers Physical Science Compatibility with Devices
 - Grade 11 Past Papers Physical Science Enhanced eBook Features
- 7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Grade 11 Past Papers Physical Science
 - Highlighting and Note-Taking Grade 11 Past Papers Physical Science
 - Interactive Elements Grade 11 Past Papers Physical Science
- 8. Staying Engaged with Grade 11 Past Papers Physical Science
 - o Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Grade 11 Past Papers Physical Science
- 9. Balancing eBooks and Physical Books Grade 11 Past Papers Physical Science
 - Benefits of a Digital Library
 - o Creating a Diverse Reading Collection Grade 11 Past Papers Physical Science
- 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
- 11. Cultivating a Reading Routine Grade 11 Past Papers Physical Science
 - Setting Reading Goals Grade 11 Past Papers Physical Science
 - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Grade 11 Past Papers Physical Science

- Fact-Checking eBook Content of Grade 11 Past Papers Physical Science
- Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
- 14. Embracing eBook Trends
 - Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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