

Isotopes

The number of protons in a nucleus determines the identity of the element. For example, any atom having 6 protons will be a "carbon" atom. If we were to add an extra proton to the nucleus, we would have an entirely different element. For example,



On the other hand, if we add an extra NEUTRON to a nucleus we simply end up with the same element, just a little heavier, since the charge on the nucleus would be unchanged.

ISOTOPES of a given element have the same ATOMIC NUMBER but a *different* ATOMIC MASS.

In other words, isotopes have the same number of protons but a different number of neutrons. An isotope is identified by its mass number, the sum of the protons and neutrons. The most common isotope of Carbon has a mass number of 12 and can be written as Carbon-12, two other isotopes are Carbon-13 and Carbon-14. Despite their different mass numbers, all three carbon isotopes react the same way chemically.

PART I. Answer the questions based on the above reading.

1. What is an isotope? **Isotopes are versions of the same element. They have the same number of protons and electrons as the element but different mass numbers and number of neutrons.**
2. What does the number next to isotopes signify? **The number indicates the isotope's mass number.**
3. How can you tell isotopes of the same element apart? **They will have a different mass number and different number of neutrons.**

PART II. For each of the following isotopes, write the number of protons, neutrons, and electrons. Assume all atoms are neutral.

	Chromium-58	Chromium-63
# of protons	24	24
# of neutrons	34	39
# of electrons	24	24

	Carbon-12	Carbon-13	Carbon-14
# of protons	6	6	6
# of neutrons	6	7	8
# of electrons	6	6	6

Isotopes Pogil Answer Key

Jiyuan Zhang

A red circular graphic with a gradient, appearing as a stylized arrow or a partial circle, located to the right of the author's name.

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