

- (ii) A compound P is heated in a test tube with sodium hydroxide solution. A red litmus paper held at the mouth of the test tube turns blue.

Which of the following could compound P be?

- (a) zinc sulphate
- (b) copper sulphate
- (c) ferrous sulphate
- (d) ammonium sulphate

- (iii) The atomic masses of sulphur (S), oxygen (O), and helium (He) are approximately 32, 16, and 4 respectively.

Which of the following statements regarding the number of atoms in 32 g of sulphur, 16 g of oxygen, and 4 g of helium is correct?

P. 16 g of oxygen contains four times the number of atoms as 4 g of helium.

Q. 16 g of oxygen contains half the number of atoms as 32 g of sulphur.

- (a) only P
- (b) only Q
- (c) both P and Q
- (d) neither P nor Q

- (iv) Ammonia gas is passed through quicklime and then collected in a jar. Red and blue litmus papers are placed in the jar. W, X, Y and Z are the four observations.

Which of the above observations correctly shows the reaction of the litmus papers to ammonia?

	Red litmus paper	Blue litmus paper
W	turns blue	remains blue
X	remains red	remains blue
Y	remains red	turns red
Z	turns blue	turns red

- (a) W
- (b) X
- (c) Y
- (d) Z

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 - The Rise of Digital Reading Icse 2chemistry Paper
 - Advantages of eBooks Over Traditional Books
2. Identifying Icse 2chemistry Paper
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Icse 2chemistry Paper
 - User-Friendly Interface
4. Exploring eBook Recommendations from Icse 2chemistry Paper
 - Personalized Recommendations
 - Icse 2chemistry Paper User Reviews and Ratings
 - Icse 2chemistry Paper and Bestseller Lists
5. Accessing Icse 2chemistry Paper Free and Paid eBooks
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 - Icse 2chemistry Paper eBook Subscription Services
 - Icse 2chemistry Paper Budget-Friendly Options

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 - Icse 2chemistry Paper Enhanced eBook Features
 7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Icse 2chemistry Paper
 - Highlighting and Note-Taking Icse 2chemistry Paper
 - Interactive Elements Icse 2chemistry Paper
 8. Staying Engaged with Icse 2chemistry Paper
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Icse 2chemistry Paper
 9. Balancing eBooks and Physical Books Icse 2chemistry Paper
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Icse 2chemistry Paper
 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
 11. Cultivating a Reading Routine Icse 2chemistry Paper
 - Setting Reading Goals Icse 2chemistry Paper
 - Carving Out Dedicated Reading Time
 12. Sourcing Reliable Information of Icse 2chemistry Paper
 - Fact-Checking eBook Content of Icse 2chemistry Paper
 - Distinguishing Credible Sources
 13. Promoting Lifelong Learning
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