

CALCULATING THE AMOUNT OF ENERGY REQUIRED TO HEAT A SPECIFIED MASS OF A SUBSTANCE USING SPECIFIC HEAT CAPACITY

Calculate the heat energy, in kJ, needed to increase the temperature of 200 g of water by 15.0°C.

Thinking	Working
Find the specific heat capacity (C) of the substance from Table 12.2.1.	The specific heat capacity of water is $4.18 \text{ J g}^{-1} \text{ }^{\circ}\text{C}^{-1}$.
To calculate the quantity of heat energy in joules, use the formula: $q = C \times m \times \Delta T$	$q = 4.18 \times 200 \times 15.0$ $= 1.25 \times 10^4 \text{ J}$
Express the quantity of energy in kJ. Remember that to convert from J to kJ, you multiply by 10^{-3} .	$q = 1.25 \times 10^4 \times 10^{-3}$ $= 12.5 \text{ kJ}$

Worked example: Try yourself 12.2.1

CALCULATING THE AMOUNT OF ENERGY REQUIRED TO HEAT A SPECIFIED MASS OF A SUBSTANCE USING SPECIFIC HEAT CAPACITY

Calculate the heat energy, in kJ, needed to increase the temperature of 375 g of water by 45.0°C.

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Whispering the Secrets of Language: An Emotional Quest through **Harcourt Achieve Unit 2 Chemistry Answers**

In a digitally-driven earth where screens reign great and quick communication drowns out the subtleties of language, the profound strategies and emotional subtleties hidden within words usually go unheard. Yet, nestled within the pages of **Harcourt Achieve Unit 2 Chemistry Answers** a charming fictional treasure pulsing with raw feelings, lies an extraordinary quest waiting to be undertaken. Written by a skilled wordsmith, that wonderful opus attracts viewers on an introspective journey, delicately unraveling the veiled truths and profound affect resonating within ab muscles cloth of each and every word. Within the emotional depths with this moving evaluation, we shall embark upon a sincere exploration of the book is key styles, dissect its fascinating writing design, and succumb to the strong resonance it evokes strong within the recesses of readers hearts.

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