

Time: 2 Hrs.

Class: XI  
Subject: Chemistry (Set - B)

M.M:50

**General Instructions:**

1. All questions are compulsory.
2. Questions 1-8 are of 1 marks each.
3. Questions 9-20 are of 2 marks each.
4. Questions 21-26 are of 3 marks each.

**Questions 1-8 are of 1 marks each:**

- Q1. Find out the volume of 0.2 mole of  $\text{SO}_2$  at S. T. P.
- Q2. How many quantum numbers are needed to designate an orbital? Name them.
- Q3. How would you justify the presence of 18 elements in the fifth period of the periodic table?
- Q4. Predict the period and the group to which an element with atomic number 40 belong to?
- Q5. Find the molarity of  $\text{AlCl}_3$  if its normality is 0.2 N.
- Q6. Draw the Lewis structure of  $\text{CO}_2$  molecule.
- Q7. How many electrons in an atom may have the following quantum number?  
 $n = 4, m_s = -\frac{1}{2}$
- Q8. What do you understand by limiting reagent? Give one use.

**Questions 9-20 are of 2 marks each**

- Q9. Sulphur and oxygen are known to form oxide. The sulphur content in one of these is 50 % while in the other is 40 %. Show that it illustrates the law of multiple proportions.
- Q10. Explain with the help of Heisenberg's Uncertainty Principle, why an electron cannot be present in the nucleus?
- Q11. The ionic size of  $\text{Cl}^-$  is greater than that of  $\text{K}^+$  though the two ions are isoelectronic. Give reason.
- Q12. Name the element which shows diagonal relationship with Li. Give reason, why it shows the diagonal relationship?
- Q13. Chlorine is prepared according to the reaction:  
 $4\text{HCl (g)} + \text{MnO}_2 \longrightarrow 2\text{H}_2\text{O} + \text{MnCl}_2 + \text{Cl}_2$   
How many grams of HCl react with 5 g of manganese dioxide? (At. Mass of Mn = 55)
- Q14. The electron energy in hydrogen atom in the first orbit is given by  $-2.18 \times 10^{-18} \text{ J atom}^{-1}$ . Calculate the energy required to remove an electron completely from the  $n = 2$  orbit. What is the longest wavelength of light that can be used to cause this transition?
- Q15. Would you expect the first ionization enthalpies of two isotopes of the same element to be same or different? Justify your statement.
- Q16. Explain the geometry of  $\text{H}_2\text{O}$  with the help of hybridization.
- Q17. Arrange the following in the increasing order of their bond length with the help of M.O. Theory:  
 $\text{O}_2, \text{O}_2, \text{O}_2^+$
- Q18. Although both  $\text{CO}_2$  and  $\text{H}_2\text{O}$  are triatomic molecules, the shape of  $\text{H}_2\text{O}$  molecule is bent while that of  $\text{CO}_2$  is linear. Explain this on the basis of dipole moment.
- Q19. Calculate molality and molarity of 93 %  $\text{H}_2\text{SO}_4$  (W/V). The density of the solution is 1.84 g/ml.
- Q20. In a hydrogen atom, an electron jumps from fourth orbit to first orbit. Find out the frequency of the spectral line using Rydberg's formula.

**Questions 21-26 are of 3 marks each.**

- Q21. What is electron gain enthalpy? Explain the factors responsible for it.
- Q22. Account for the following:
- a) Among hydrides of group 16 elements, which has the lowest boiling point? Give reason.
  - b) Write the differences between BMO and ABMO. (Two)
  - c) Draw the resonating structures of  $\text{NO}_2^-$ .

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