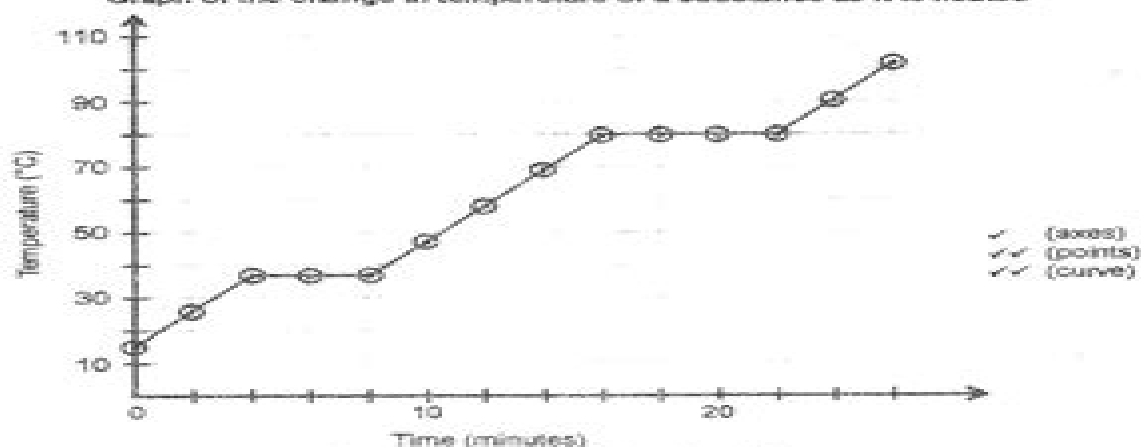


3.4

Graph of the change in temperature of a substance as it is heated

(5)



3.5 Substance exists as both a solid and a liquid ✓✓ or/... is melting (2)

3.6 37 °C ✓ (2)

3.7 80 °C ✓ (1)

3.8 An indication of the average kinetic energy ✓ of the particles in a substance. ✓ (2)

3.9 Physical ✓ (1)

3.10 In the solid state the intermolecular forces are the strongest and the particles cannot move. The heater provides thermal energy to the system, which is translated into kinetic energy (increase in temperature) until the intermolecular forces between the particles can be broken. ✓✓ At this point the thermal energy is transformed into latent energy to break the bonds, so no increase in temperature occurs. ✓✓ (4)

[23]

QUESTION 4

4.1

4.1.1 Homogeneous ✓ (1)

4.1.2 Homogeneous ✓ (1)

4.1.3 Pure substance ✓ (1)

4.2

4.2.1 Potassium bromide ✓ (2)

4.2.2 Calcium carbonate ✓ (2)

4.2.3 Ammonium chloride ✓ (2)

4.3

4.3.1 ZnSO_4 ✓✓ (2)4.3.2 KMnO_4 ✓✓ (2)4.3.3 H_2O ✓✓ (2)

[17]

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