

Module 4 Practice Problems

1. Nitric acid is produced with the help of nitrous oxide. One disadvantage of N_2O will be discussed because this is a greenhouse contribution.



$$\frac{1 \text{ mol } \text{N}_2\text{O}}{1 \text{ mol } \text{C}_2\text{H}_6} \quad 0.17 \text{ mol } \text{C}_2\text{H}_6 \cdot \left(\frac{1 \text{ mol } \text{N}_2\text{O}}{1 \text{ mol } \text{C}_2\text{H}_6} \right) = \boxed{0.17 \text{ mol } \text{N}_2\text{O}}$$

2. Some waste water treatment plants have (all kinds) of anaerobic bacteria on the bed produced about 0.200 mol of methane per day. ~~1.0000 or 1.00 mol~~



3. From a very small amount of H_2O is produced in the following reaction:



How much water is produced from 1.0 g of H_2O_2 ? (Molar mass of H_2O_2 is 34.01 g/mol, molar mass of H_2O is 18.02 g/mol)

$$\frac{5 \text{ mol } \text{H}_2\text{O}}{1 \text{ mol } \text{H}_2\text{O}_2} \quad 1 \text{ g } \text{H}_2\text{O}_2 \cdot \left(\frac{5 \text{ mol } \text{H}_2\text{O}}{1 \text{ mol } \text{H}_2\text{O}_2} \right) = \boxed{0.294 \text{ mol } \text{H}_2\text{O}}$$

4. The decomposition of hydrogen peroxide is represented by the following reaction: ~~1.0000 or 1.00 mol~~



$$1 \text{ mol } \text{H}_2\text{O}_2 \cdot \left(\frac{1 \text{ mol } \text{O}_2}{2 \text{ mol } \text{H}_2\text{O}_2} \right) = \boxed{0.5 \text{ mol } \text{O}_2}$$

5. The reaction of phosphorus trichloride is represented by the following reaction:



How much H_3PO_3 is produced from 1.00 g of PCl_3 ? (Molar mass of PCl_3 is 137.33 g/mol, molar mass of H_3PO_3 is 81.99 g/mol)

$$1 \text{ mol } \text{PCl}_3 \cdot \left(\frac{1 \text{ mol } \text{H}_3\text{PO}_3}{1 \text{ mol } \text{PCl}_3} \right) = 1 \text{ mol } \text{H}_3\text{PO}_3 \cdot \left(\frac{81.99 \text{ g}}{1 \text{ mol}} \right) = \boxed{81.99 \text{ g } \text{H}_3\text{PO}_3}$$

6. From decomposition of calcium carbonate, the product carbon dioxide is decomposed ~~1.0000 or 1.00 mol~~



$$1 \text{ mol } \text{CaCO}_3 \cdot \left(\frac{1 \text{ mol } \text{CO}_2}{1 \text{ mol } \text{CaCO}_3} \right) = 1 \text{ mol } \text{CO}_2 \cdot \left(\frac{44.01 \text{ g}}{1 \text{ mol}} \right) = \boxed{44.01 \text{ g } \text{CO}_2}$$

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M Woodhall



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