MOLE RATIOS IN CHEMICAL EQUATIONS

When the equation is balanced, you can use the coefficients to get the mole ratios between all substances present to get the amount (moles) of each (pure) substance.

Remember, that one mole of a substance is equal to Avogadro's number $(N_4) = 6.022 \times 10^{23} \text{ particles/mol}$ of particles or entities of that substance (which could be a reactant of product). Below is a list of what an 'entity' represents for each type of pure substance.

Element - atom e.g. Al Covalent/Molecular Compound - molecule e.g. H,SO, Ionic Compound - formula unit e.g. Al₂(SO₄)₃

There is more than one way to set up/look at a mole ratio, but ultimately the final answer should be the same. There are short cuts too, but you should only use these after you understand what is happening.

MOLE RATIOS

"cross multiplication" method

so, if you have 5.00 moles of Al, then how many moles of H₂SO₄ are required to react with the Al?

G x R/G method

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