

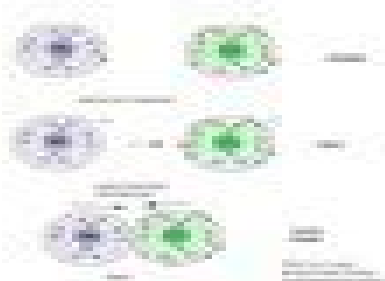
## Final Study Guide

**Valence Electrons:** Electrons on the outermost shell or orbital, electrons involved in bonding, atoms are most stable when they have 8 (this is called the octet rule)

**Isotopes:** atoms with the same atomic number, but different atomic masses, because they have different amount of neutrons

**Oxidation Numbers:** positive or negative number shows the symbol that tells us how many electrons an atom has to lose or gain

**Ionic Bond:** a bond between a metal and a non-metal where electrons are lost and gained



Metal usually loses the electron and the nonmetal usually gains it, this causes the metal to be positively charged and the nonmetal to be negatively charged. Positives and negatives attract.

**Properties:**

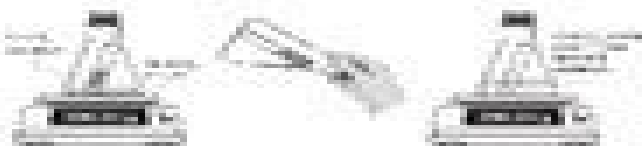
- Conduct electricity when melted or dissolved in water.
- High melting points
- Are polar
- Usually form crystalline structures

**Covalent Bonds:** a bond between two nonmetals where electrons are shared, both nuclei are attracted to the shared electrons. Electrons can be shared, unequally which results in a polar covalent bond. Can also form double and triple bonds.

**Properties:**

- Low melting points
- Tend to be softer
- Usually don't dissolve in water
- Don't conduct electricity

**The Law of Conservation of mass:** Mass cannot be created or destroyed. The mass of the products must equal the mass of the reactants



as you can see in this picture the mass before the reaction is equal to the mass after the reaction.

This rule applies to atoms as well.

You must have the same number and type of atoms at the beginning of the reaction as you do after the reaction

### Steps for Balancing Equations

1. List the atoms involved in the reaction
2. Count how many of each type of atom on each side
3. Add coefficients to make the number of each type of atom on each side be equal
4. Recount the number of each type of atom on each side
5. If not equal repeat steps 3-4

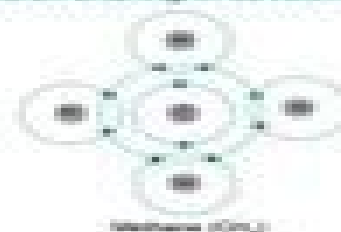
**Structure of water molecule**



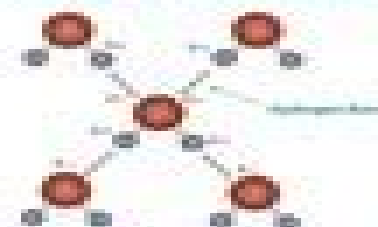
**Water**

- Is a polar molecule

**Electron sharing in covalent bond**



**Methane (CH<sub>4</sub>)**



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