



CHEMICAL BONDS



IONIC, COVALENT, & METALLIC

FORMS OF CHEMICAL BONDS
Bonding: combination of atoms

Atoms combine to form molecules or compounds by sharing or transferring electrons. This process is called bonding. The resulting molecule or compound has properties different from the individual atoms.

Types of Bonds:

- Covalent Bond:** Formed by sharing electrons between two atoms. Example: H_2 , H_2O , CO_2 .
- Ionic Bond:** Formed by the transfer of electrons from one atom to another, creating oppositely charged ions that attract each other. Example: $NaCl$, MgO .
- Metallic Bond:** Formed by a sea of delocalized electrons surrounding positive metal ions. Example: Fe , Cu , Al .

Covalent, Ionic, and Metallic Bonding

Covalent bonding occurs when two atoms share one or more pairs of electrons. This results in a stable molecule. Ionic bonding occurs when one atom transfers an electron to another atom, creating a positive ion and a negative ion. These ions are attracted to each other. Metallic bonding occurs in metals, where the atoms are packed closely together, and the outer electrons are delocalized, forming a 'sea of electrons' that holds the metal together.

Science Real




Diagram illustrating the formation of chemical bonds:

The diagram shows the interaction of two atoms, each with a nucleus (positive charge) and a valence shell (outermost shell). The atoms are shown with their respective valence electrons (represented by dots).

Top Right: A diagram showing the formation of an ionic bond. One atom (red nucleus, negative charge) loses an electron, becoming a cation (positive charge). The other atom (yellow nucleus, positive charge) gains an electron, becoming an anion (negative charge). The resulting ionic bond is shown as a pair of oppositely charged ions.

Bottom Right: A diagram showing the formation of a covalent bond. Two atoms (yellow nuclei, positive charges) share a pair of electrons, forming a covalent bond. The resulting molecule is shown as two atoms sharing a pair of electrons.

Bottom Left: A diagram showing the formation of a metallic bond. A cluster of atoms (yellow nuclei, positive charges) is shown with a sea of delocalized electrons (represented by a blue cloud) surrounding them. The resulting metallic bond is shown as a cluster of atoms held together by the sea of electrons.

Bottom Center: A diagram showing the formation of a double bond. Two atoms (yellow nuclei, positive charges) share two pairs of electrons, forming a double bond. The resulting molecule is shown as two atoms sharing two pairs of electrons.

Bottom Right (Tablet): A diagram showing the formation of a triple bond. Two atoms (yellow nuclei, positive charges) share three pairs of electrons, forming a triple bond. The resulting molecule is shown as two atoms sharing three pairs of electrons.

Bottom Right (Text): A text box explaining the difference between single, double, and triple covalent bonds. It states that single bonds are the weakest, double bonds are stronger, and triple bonds are the strongest.

GUIDED READING ACTIVITY

Ionic Bonds Guided Reading And Study Answers

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