

- Directions:**
 1. An element may be assigned a positive or negative oxidation number of zero.
 2. An atom with a higher electronegativity is assigned a negative oxidation number.
 3. The sum of all oxidation numbers for a molecule must be zero.

Ion Practice Set

1. Write an equation that shows how the element is oxidized or reduced.
 2. Write the net ionic equation for the reaction.
 3. Write the net ionic equation for the reaction.

Complete the following table, using the periodic table in the back of your book.

	ELEMENT NAME	ION SYMBOL	NUMBER OF PROTONS	NUMBER OF ELECTRONS	CHARGE OF ION	TYPE OF ION
1	Sodium	Na^+	11	10	+1	cation
2	Chlorine	Cl^-	17	18	-1	anion
3	Strontium	Sr^{2+}	38	36	+2	cation
4	Antimony	Sb^{3-}	51	54	-3	anion
5	Calcium	Ca^{2+}	20	18	+2	cation
6	Barium	Ba^{2+}	56	54	+2	cation
7	Hydrogen	H^+	1	0	+1	cation
8	Lithium	Li^+	3	2	+1	cation
9	Aluminum	Al^{3+}	13	10	+3	cation
10	Sulfur	S^{2-}	16	18	-2	anion
11	Strontium	Sr^{2+}	38	36	+2	cation
12	Hydrogen	H^+	1	0	+1	cation
13	Lithium	Li^+	3	2	+1	cation
14	Strontium	Sr^{2+}	38	36	+2	cation
15	Strontium	Sr^{2+}	38	36	+2	cation

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DP Hallahan

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Table of Contents Ion Practice Set Og Answers

1. Understanding the eBook Ion Practice Set Og Answers
 - The Rise of Digital Reading Ion Practice Set Og Answers
 - Advantages of eBooks Over Traditional Books
2. Identifying Ion Practice Set Og Answers
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Ion Practice Set Og Answers
 - User-Friendly Interface
4. Exploring eBook Recommendations from Ion Practice Set Og Answers
 - Personalized Recommendations
 - Ion Practice Set Og Answers User Reviews and Ratings
 - Ion Practice Set Og Answers and Bestseller Lists
5. Accessing Ion Practice Set Og Answers Free and Paid eBooks
 - Ion Practice Set Og Answers Public Domain eBooks
 - Ion Practice Set Og Answers eBook Subscription Services

- Ion Practice Set Og Answers Budget-Friendly Options
- 6. Navigating Ion Practice Set Og Answers eBook Formats
 - ePub, PDF, MOBI, and More
 - Ion Practice Set Og Answers Compatibility with Devices
 - Ion Practice Set Og Answers Enhanced eBook Features
- 7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Ion Practice Set Og Answers
 - Highlighting and Note-Taking Ion Practice Set Og Answers
 - Interactive Elements Ion Practice Set Og Answers
- 8. Staying Engaged with Ion Practice Set Og Answers
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Ion Practice Set Og Answers
- 9. Balancing eBooks and Physical Books Ion Practice Set Og Answers
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Ion Practice Set Og Answers
- 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
- 11. Cultivating a Reading Routine Ion Practice Set Og Answers
 - Setting Reading Goals Ion Practice Set Og Answers
 - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Ion Practice Set Og Answers
 - Fact-Checking eBook Content of Ion Practice Set Og Answers
 - Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
- 14. Embracing eBook Trends

- Integration of Multimedia Elements
- Interactive and Gamified eBooks

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