

- Directions:**
 1. An element may be assigned a positive or negative oxidation number.
 2. An atom with a higher electronegativity is assigned a negative oxidation number.
 3. The sum of all oxidation numbers for a molecule must be zero.

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1. Write an equation that shows how the element is oxidized or reduced.
 2. Write the oxidation number next to the element.
 3. Write the oxidation number next to the element.

Complete the following table, using the periodic table in the back of your book.

	ELEMENT NAME	ION SYMBOL	NUMBER OF PROTONS	NUMBER OF ELECTRONS	NUMBER OF ELECTRONS LOST OR GAINED	TYPE OF ION
1	Sodium	Na^+	11	10	lost one	cation
2	Chlorine	Cl^-	17	18	gained one	anion
3	Sulfur	S^{2-}	16	18	gained two	anion
4	Calcium	Ca^{2+}	20	18	lost two	cation
5	Carbon	C^{4-}	6	10	gained four	anion
6	Aluminum	Al^{3+}	13	10	lost three	cation
7	Hydrogen	H^+	1	0	lost one	cation
8	Oxygen	O^{2-}	8	10	gained two	anion
9	Magnesium	Mg^{2+}	12	10	lost two	cation
10	Fluorine	F^-	9	10	gained one	anion
11	Silver	Ag^+	47	46	lost one	cation
12	Strontium	Sr^{2+}	38	36	lost two	cation
13	Barium	Ba^{2+}	56	54	lost two	cation
14	Lead	Pb^{2+}	82	80	lost two	cation
15	Copper	Cu^+	29	28	lost one	cation
16	Iron	Fe^{2+}	26	24	lost two	cation
17	Zinc	Zn^{2+}	30	28	lost two	cation
18	Gold	Au^+	79	78	lost one	cation

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Robert R. Redfield



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