# **Chemistry 100 Final Study Guide**

#### Chapter 1: Introduction

- a. Physical vs. Chemical properties
  - 1) List 2 physical and 1 chemical property of magnesium.
- b. Units of measurement
  - Can you distinguish between metric and U.S. units of measurement?
  - What is 0.000000568 in scientific notation? What is 4.63 x 107 in regular notation?
- c. Denisity
  - 1) What is the density of a 37g block of metal that takes up 15mL of volume?

#### Chapter 3: Atomic Structure

- a. Historical Perspective
  - List the major scientific accomplishments of the following scientists:
    - (i) Michael Faraday
    - (ii) Joseph Thomson
    - (iii) Robert Millikan
    - (iv) Emest Rutherford
    - (v) Neils Bohr
    - (vi) Envin Schrodinger
- b. Isotopes
  - Define an isotope and give an example of several isotopes for hydrogen.
- Structure of the atom
  - 1) Name the 3 main parts of an atom, and indicate where in the atom they'd be found.
  - 2) For the element sodium: how many energy levels? How many sublevels? What orbitals? How many electrons?
  - 3) Write the full electron configuration of silicon.
- d. Electron configurations and trends on the Periodic Table
  - Locate each of the following on the Periodic Table: alkali metals, alkaline earth metals, transition metals, halogens, noble gases, metals, non-metals.

#### Chapter 5: Chemical bonds

- a. Electron-dot structures
  - Draw the electron-dot structure of an oxygen atom.
- b. Octet Rule and lons
  - 1) What charge of ions do magnesium, chlorine and nitrogen form?
- c. Ionic compounds
  - Name the following compounds: AlBr<sub>3</sub>, CaS, Na<sub>3</sub>N
- d. Covalent compounds
  - Name the following compounds: SO<sub>2</sub>, PCl<sub>3</sub>, CH<sub>4</sub>
- e. Polyatomic molecules
  - Name and write the formula for the 6 –ate polyatomic molecules we discussed.
- Electron-dot formulas for molecules
  - 1) Write the electron-dot formula for each of the 6 -ate polyatomic ions we discussed.
- g. VSEPR Model and Molecular shape
  - Draw the electron-dot formula of ammonia (NH<sub>2</sub>) and state its VSEPR shape.
- Molecule polarity
  - 1) Which of the VSEPR molecule shapes are polar and which are non-polar?

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J Elliott

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