

1. Write the chemical equation for the reversible reaction studied in this experiment:



2. Write the equilibrium expression for the reaction studied in this experiment:

$$K_c = \frac{\text{FeSCN}^{2+}(\text{aq})}{(\text{Fe}^{3+})_{\text{aq}}(\text{SCN}^{-})_{\text{aq}}}$$

3. Complete the following table from your observations of test tubes #1-#6:

Test Tube	Change	Initial Color	Final Color (darker/lighter)
	Initial color before anything is added		
1	Added $\text{Fe}(\text{NO}_3)_3$	Reddish color	darker
2	Added KSCN	Reddish color	darker
3	Added $\text{NH}_4\text{Cl}$	Reddish color	lighter
4	Added $\text{AgNO}_3$	Reddish color	lighter
5	Cold	Reddish color	darker
6	Hot	Reddish color	lighter

4. Complete the following table for test tubes #2 - 4:

Test Tube	What concentration was changed and how it was changed?	The equilibrium shifted to left/right/no change	What evidence for the direction of shift do you have?
2	$\text{FeSCN}^{2+}$ is a mix of an iron/reddish color + thiocyanate has no color	Shifted Right	The Le Chatelier principle w/ reaction from left to right occurs when more reactant was added.
3	$[\text{SCN}^{-}]$ was increased	Shifted Right	As more reactant was added, the concentration of the product goes up
4	$\text{Fe}^{3+}$ concentration was changed	Shifted Left	$\text{FeCO}_3$ will form which leads to a lower $\text{Fe}^{3+}$ concentration

# Lab 1rate And Equilibrium Report Answer

**DJ Losen**



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## **Table of Contents Lab 1rate And Equilibrium Report Answer**

1. Understanding the eBook Lab 1rate And Equilibrium Report Answer
  - The Rise of Digital Reading Lab 1rate And Equilibrium Report Answer
  - Advantages of eBooks Over Traditional Books
2. Identifying Lab 1rate And Equilibrium Report Answer
  - Exploring Different Genres
  - Considering Fiction vs. Non-Fiction
  - Determining Your Reading Goals
3. Choosing the Right eBook Platform
  - Popular eBook Platforms
  - Features to Look for in an Lab 1rate And Equilibrium Report Answer
  - User-Friendly Interface
4. Exploring eBook Recommendations from Lab 1rate And Equilibrium Report Answer
  - Personalized Recommendations
  - Lab 1rate And Equilibrium Report Answer User Reviews and Ratings
  - Lab 1rate And Equilibrium Report Answer and Bestseller Lists
5. Accessing Lab 1rate And Equilibrium Report Answer Free and Paid eBooks
  - Lab 1rate And Equilibrium Report Answer Public Domain eBooks
  - Lab 1rate And Equilibrium Report Answer eBook Subscription Services
  - Lab 1rate And Equilibrium Report Answer Budget-Friendly Options

6. Navigating Lab 1rate And Equilibrium Report Answer eBook Formats
  - ePub, PDF, MOBI, and More
  - Lab 1rate And Equilibrium Report Answer Compatibility with Devices
  - Lab 1rate And Equilibrium Report Answer Enhanced eBook Features
7. Enhancing Your Reading Experience
  - Adjustable Fonts and Text Sizes of Lab 1rate And Equilibrium Report Answer
  - Highlighting and Note-Taking Lab 1rate And Equilibrium Report Answer
  - Interactive Elements Lab 1rate And Equilibrium Report Answer
8. Staying Engaged with Lab 1rate And Equilibrium Report Answer
  - Joining Online Reading Communities
  - Participating in Virtual Book Clubs
  - Following Authors and Publishers Lab 1rate And Equilibrium Report Answer
9. Balancing eBooks and Physical Books Lab 1rate And Equilibrium Report Answer
  - Benefits of a Digital Library
  - Creating a Diverse Reading Collection Lab 1rate And Equilibrium Report Answer
10. Overcoming Reading Challenges
  - Dealing with Digital Eye Strain
  - Minimizing Distractions
  - Managing Screen Time
11. Cultivating a Reading Routine Lab 1rate And Equilibrium Report Answer
  - Setting Reading Goals Lab 1rate And Equilibrium Report Answer
  - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Lab 1rate And Equilibrium Report Answer
  - Fact-Checking eBook Content of Lab 1rate And Equilibrium Report Answer
  - Distinguishing Credible Sources
13. Promoting Lifelong Learning
  - Utilizing eBooks for Skill Development
  - Exploring Educational eBooks
14. Embracing eBook Trends
  - Integration of Multimedia Elements

- Interactive and Gamified eBooks

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