

CHAPTER 7

STUDY GUIDE

Ionic Compounds and Metals

Section 7.1 Ion Formation

In your textbook, read about chemical bonds and formation of ions.

Use each of the terms below just once to complete the passage.

chemical bond	electrons	energy level	ions	noble gases
nucleus	octet	pseudo-noble gas formations	valence	

The force that holds two atoms together is called a(n) (1) _____.
Such an attachment may form by the attraction of the positively charged
(2) _____ of one atom for the negatively charged
(3) _____ of another atom, or by the attraction of charged atoms, which are called
(4) _____. The attractions may also involve
(5) _____ electrons, which are the electrons in the outermost
(6) _____. The (7) _____ are a family of elements that have very
little tendency to react. Most of these elements have a set of eight outermost electrons, which is called a
stable (8) _____. The relatively stable electron structures developed by loss of
electrons in certain elements of groups 3, 4, 13, and 14 are called (9) _____.

For each statement below, write *true* or *false*.

- _____ 10. A positively charged ion is called an anion.
- _____ 11. Elements in group 1 lose their one valence electron, forming an ion with a 1+ charge.
- _____ 12. Elements tend to react so that they acquire the electron structure of a halogen.
- _____ 13. A sodium atom tends to lose one electron when it reacts.
- _____ 14. The electron structure of a zinc ion (Zn^{2+}) is an example of a pseudo-noble gas formation.
- _____ 15. A Cl^{-} ion is an example of a cation.
- _____ 16. The ending *-ide* is used to designate an anion.
- _____ 17. Nonmetals form a stable outer electron configuration by losing electrons and becoming anions.

Ionic Compounds And Metals Study Guide

Leo J. Malone, Theodore O. Dolter



Ionic Compounds And Metals Study Guide:

Chemistry, Student Study Guide James E. Brady, Fred Senese, 2008-01-28 The image on the front cover depicts a carbon nanotube emerging from a glowing plasma of hydrogen and carbon as it forms around particles of a metal catalyst Carbon nanotubes are a recently discovered allotrope of carbon Three other allotropes of carbon buckyballs graphite and diamond are illustrated at the left as is the molecule methane CH_4 from which nanotubes and buckyballs can be made The element carbon forms an amazing number of compounds with structures that follow from simple methane found in natural gas to the complex macromolecules that serve as the basis of life on our planet The study of chemistry also follows from the simple to the more complex and the strength of this text is that it enables students with varied backgrounds to proceed together to significant levels of achievement

A-Level Study Guide Chemistry Ed H2.2 CS Toh, 2016-03-08 This is an ebook version of the A Level Study Guide Chemistry Higher 2 Ed H2 2 published by Step by Step International Pte Ltd For the revised Higher 2 H2 syllabus with first exam in 2017 This ebook gives concise illustrated notes and worked examples It is intended as a study guide for readers who have studied the O Level Chemistry or the equivalent It contains material that most readers should want to take note of when attending formal lessons and or discussions on the Singapore Cambridge GCE A Level Higher 2 H2 Chemistry As the Higher 1 H1 Chemistry syllabus is a subset of the H2 Chemistry syllabus this ebook is also suitable for readers studying Chemistry at the H1 level The concise notes cover essential steps to understand the relevant theories The illustrations and worked examples show essential workings to apply those theories We believe the notes and illustrations will help readers learn to learn and apply the relevant knowledge The ebook should help readers study and prepare for their exams Relevant feedbacks from Examiner Reports reflecting what the examiners expected are incorporated into the notes and illustrations where possible or appended as notes NB where appropriate It is also a suitable aid for teaching and revision

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Odian,2006-01-27 This study guide provides reader friendly reinforcement of the concepts covered in the textbook Features include Chapter outlines Are you able to Worked text problems Fill ins Test yourself Concept maps Can also be used for Blei and Odian s Organic and Biochemistry **Study Guide to Accompany Basics for Chemistry** Martha Mackin,2012-12-02 Study Guide to Accompany Basics for Chemistry is an 18 chapter text designed to be used with Basics for Chemistry textbook Each chapter contains Overview Topical Outline Skills and Common Mistakes which are all keyed to the textbook for easy cross reference The Overview section summarizes the content of the chapter and includes a comprehensive listing of terms a summary of general concepts and a list of numerical exercises while the Topical Outline provides the subtopic heads that carry the corresponding chapter and section numbers as they appear in the textbook The Fill in Multiple Choice are two sets of questions that include every concept and numerical exercise introduced in the chapter and the Skills section provides developed exercises to apply the new concepts in the chapter to particular examples The Common Mistakes section is designed to help avoid some of the errors that students make in their effort to learn chemistry while the Practical Test section includes matching and multiple choice questions that comprehensively cover almost every concept and numerical problem in the chapter After briefly dealing with an overview of chemistry this book goes on exploring the concept of matter energy measurement problem solving atom periodic table and chemical bonding These topics are followed by discussions on writing names and formulas of compounds chemical formulas and the mole chemical reactions calculations based on equations gases and the properties of a liquid The remaining chapters examine the solutions acids bases salts oxidation reduction reactions electrochemistry chemical kinetics and equilibrium and nuclear organic and biological chemistry This study guide will be of great value to chemistry teachers and students **Basic Concepts of Chemistry, 9e Study Guide and Solutions Manual** Leo J. Malone,Theodore O. Dolter,2012-01-03 The 9th edition of Malone s Basic Concepts of Chemistry provides many new and advanced features that continue to address general chemistry topics with an emphasis on outcomes assessment New and advanced features include an objectives grid at the end of each chapter which ties the objectives to examples within the sections assessment exercises at the end each section and relevant chapter problems at the end of each chapter A new Math Check allows quick access to the needed basic skill The first chapter now includes brief introductions to several fundamental chemical concepts and Chapter Synthesis Problems have been added to the end of each chapter to bring key concepts into one encompassing problem Every concept in the text is clearly illustrated with one or more step by step examples Making it Real essays have been updated to present timely and engaging real world applications emphasizing the relevance of the material they are learning This edition continues the end of chapter Student Workshop activities to cater to the many different learning styles and to engage users in the practical aspect of the material discussed in the chapter *Chemical Principles Study Guide/Solutions Manual* John Krenos,Joseph Potenza,Carl Hoeger,2007-01-18 Written for general chemistry courses Chemical Principles helps students develop chemical insight by showing the

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