

Electron Configuration of Ions

Write out the configuration of each ion on the line. Then use the key in the right column to match the configuration to the noble gas that has the same configuration. Most of the noble gases will be used more than once. Some may not be used at all.

1. Na^+ : _____2. F^- : _____3. K^+ : _____4. Sr^{2+} : _____5. Mg^{2+} : _____6. O^{2-} : _____7. Br^- : _____8. P^{3-} : _____9. Li^+ : _____10. Fe^{2+} : _____11. S^{2-} : _____12. Ca^{2+} : _____13. Cl^- : _____14. Zn^{2+} : _____15. Sc^{3+} : _____

A. Helium (He)

B. Neon (Ne)

C. Argon (Ar)

D. Krypton (Kr)

E. Xenon (Xe)

F. Radon (Rn)

G. Does not match a noble gas.

He

Ne

Ar

Kr

Xe

Rn

Ion Electron Configuration Worksheet 2014

Dale Rich



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Electron Configuration 148 Success Secrets - 148 Most Asked Questions on Electron Configuration - What You Need to Know Dale Rich, 2014-10-15 It's a brand new Electron configuration world There has never been a Electron configuration Guide like this It contains 148 answers much more than you can imagine comprehensive answers and extensive details and references with insights that have never before been offered in print Get the information you need fast This all embracing guide offers a thorough view of key knowledge and detailed insight This Guide introduces what you want to know about Electron configuration A quick look inside of some of the subjects covered Extended periodic table 8s elements List of data references for chemical elements References for chemical elements Ferrocene Structure and bonding Coordination complex Geometry History of electromagnetic theory Lorentz and Poincare Ion Formation of polyatomic and molecular ions Valence electrons Copernicium Physical and atomic Cadmium Mol Phys Term symbol Valence electron Electron configuration Pnictogen Chemical Electronic configuration Electron configuration in molecules Group 11 elements Hund's rules Octet rule Explanation in quantum theory Group 11 element Characteristics Hassium Physical and atomic Alkali metal Extended periodic table Superactinides Neutron absorption Neutron absorbers Flerovium Nuclear stability and isotopes Lewis acid Comparison with Bronsted Lowry Theory Ununoctium Calculated atomic and physical properties Tunnel diode f block Isotopes of strontium Ytterbium Aufbau principle Atoms Energy levels Pauli exclusion principle Atoms and the Pauli principle Electron shell List of elements with electrons per shell Periodic table Layout Rutherfordium Chemical Antimony Group 8 element Lanthanum Chemical properties Extended periodic table Eka superactinides Manganese Promethium Physical properties and much more

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