

Molarity = Moles of solute / Liters of Solution (abbreviation = M)

Molality = Moles of solute / Kg of Solvent (abbreviation = m)

Normality = number of equivalent of solute x Molarity of Solution (abbreviation = N)

Mass Percent = mass of solute / mass of solution

- How many moles of ethyl alcohol,  $C_2H_5OH$ , are present in 65 mL of a 1.5 M solution?
- How many liters of a 6.0 M solution of acetic acid  $CH_3COOH$ , contain 0.0030 moles of acetic acid?
- You want 85 g of KOH. How much of a 3.0 m solution of KOH will provide it?
- If you dissolve 0.70 moles of HCl in enough water to prepare 250 mL of solution, what is the molarity of the solution you have prepared?
- A solution is prepared by adding 2.0 L of 6.0 M HCl to 500 mL of a 9.0 M HCl solution. What is the molarity of the new solution? (Remember, the volumes are additive)
- Convert the following Molarities to Normalities.
  - 2.5 M HCl = \_\_\_\_\_ N
  - 1.4 M  $H_2SO_4$  = \_\_\_\_\_ N
  - 1.0 M NaOH = \_\_\_\_\_ N
  - 0.5 M  $Ca(OH)_2$  = \_\_\_\_\_ N
- A commonly purchased disinfectant is a 3.0% (by mass) solution of hydrogen peroxide ( $H_2O_2$ ) in water. Assuming the density of the solution is 1.0 g/cm<sup>3</sup>, calculate the molarity and molality of  $H_2O_2$ .
- A solution is made by dissolving 25 g of NaCl in enough water to make 1.0 L of solution. Assume the density of the solution is 1.0 g/cm<sup>3</sup>. Calculate the molarity and molality of the solution.
- What is the mass percent of a solution that contains 152 g of  $KNO_3$  in 7.86 L of water?
- A solution has been prepared to be 9.00 % (by mass) glucose. For every 100.0 grams of solution present there are \_\_\_\_\_ grams of glucose.
- How many grams of KBr are contained in 250. grams of a 6.25 % KBr solution?
- What is the concentration of each type of ion and total concentration of ions in a 0.375 M Ammonium phosphate solution?
- How many moles of chloride ions are in 1.50 L of 4.15 M zinc chloride?

#### ANSWERS

- 0.098 moles of alcohol
- $5.0 \times 10^{-4}$  L solution
- $5.9 \times 10^{-2}$  g solution
- 2.8M
- 6.6 M
- 2.5 N
  - 2.8 N
  - 1.0 N
  - 1 N
- 0.88M, 0.91m
- 0.43M, 0.43m
- 1.90 % by mass of  $KNO_3$
- 9.00
- 15.6 g KBr
- 1.13 M of  $(NH_4)^{+1}$  and 0.375 M of  $(PO_4)^{-3}$
  - 1.51 M of ions
- 12.45 moles of  $Cl^{-1}$

# Holt Chemistry Section Concentration And Molarity Answers

**Jin-Ying Zhang**



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