

## The Great Gatsby Study Guide

### Chapter 5



1. Language usage: provide a definition for each of the following words.

- distraught
- vestige
- muster
- ecstatic
- nebulous
- render
- obstinate
- tactless
- counterfeit

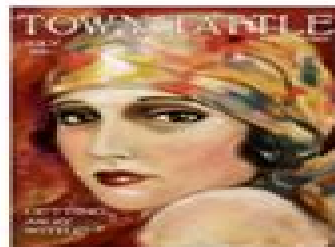
2. Briefly summarise this chapter.

3. Reading Comprehension: Answer each question in as much detail as you can. Provide enough information to establish that you read the chapter carefully.

- Why does Gatsby have "the grass cut" at the beginning of Chapter 5 (82)? Whose grass gets cut?
- What does Gatsby reply when Nick asks him how he makes his money? Why does Nick find that significant?
- Why do you think Daisy sobs when Gatsby shows her his shirts?

4. Characterisation:

1. How does Fitzgerald use indirect characterization to illustrate Gatsby's nervousness in this chapter? Provide direct evidence from the text.
2. Find three examples about Daisy's voice in this chapter. Why might Fitzgerald chose to focus on her voice as part of her representation? What might it illustrate about her?



# Great Gatsby Chapter 5 Study Guide Answers

**Wenbin Ji**



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web in the context of a chemical reaction the term neutralization is used for a reaction between an acid and a base or alkali historically this reaction was represented as acid base alkali salt water for example  $\text{hcl naoh nacl h}_2\text{o}$

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web 1 sodium hydroxide solution is poured into a beaker of hydrochloric acid which contains a thermometer showing room temperature 2 the beaker now contains sodium chloride and water and the

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web reacting two solutions eg acid and alkali place the polystyrene cup inside the glass beaker to make it more stable

measure an appropriate volume of each liquid eg 25 cm<sup>3</sup> place one of the

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web for reactions involving acetic acid or ammonia the measured enthalpy change of neutralization is a few kJ less exothermic than with strong acids and bases for example one source which gives the enthalpy change of neutralization of sodium hydroxide solution with HCl as 57.9 kJ mol<sup>-1</sup>

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web sep 20 2014 i will illustrate this with some examples when an acid is neutralised by an alkali the reaction is exothermic

eg 1 HCl(aq) + NaOH(aq) → NaCl(aq) + H<sub>2</sub>O(l) for which ΔH = 57 kJ mol<sup>-1</sup> eg 2 HNO<sub>3</sub>(aq) + KOH(aq) → KNO<sub>3</sub>(aq) + H<sub>2</sub>O(l) for which ΔH = 57 kJ mol<sup>-1</sup>

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web the enthalpy change of neutralisation is defined as the enthalpy change when solutions of an acid and alkali react together to produce one mole of water for example HCl(aq) + NaOH(aq) → NaCl(aq) + H<sub>2</sub>O(aq) you can see that one mole of water is produced when one mole of HCl or one mole of NaOH reacts

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