



Periodic Table of the Elements

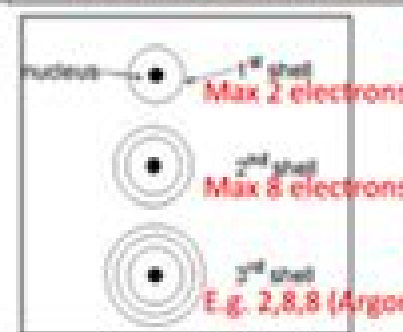
- metals
- non-metals
- transition metals
- alkali metals
- alkaline earth metals
- halogens
- noble gases
- lanthanoids
- actinoids

An atom:
 Number of Protons = Number of Electrons
 Mass number = Protons + Neutrons

C2 Chemistry

Periodic table arranges elements by **ATOMIC NUMBER** (proton number/the small one)

The number of outer shell electrons match the group the element is found in.
 E.g. Lithium 2,1 is a group 1 element.

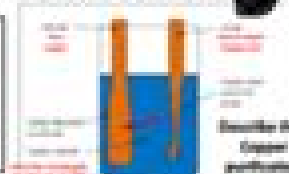


Protons and neutrons both have a mass of 1

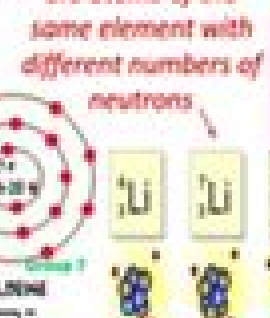
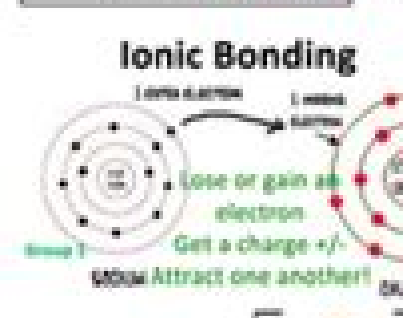
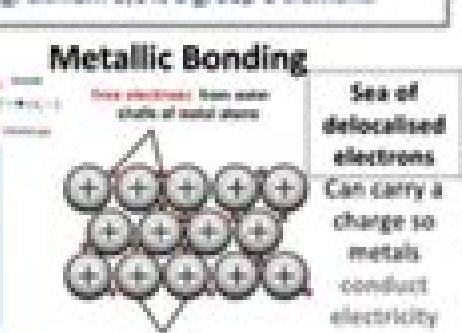
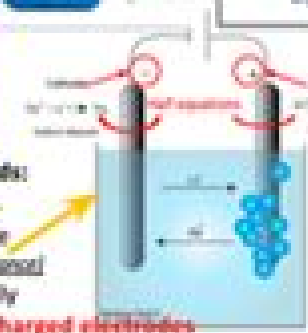
ISOTOPES are atoms of the same element with different numbers of neutrons

OIL RIG

Oxidation is loss (of electrons)
 Reduction is gain (of electrons)



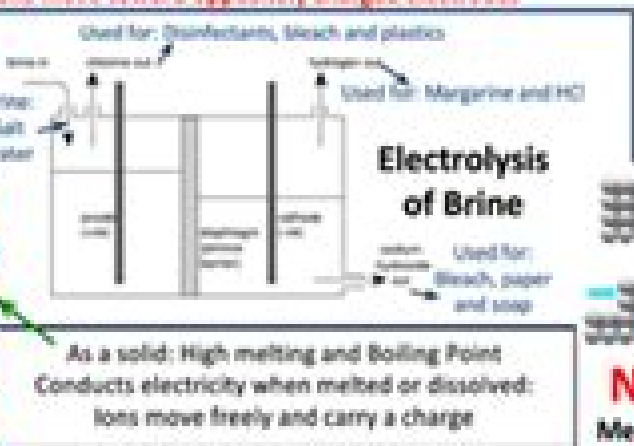
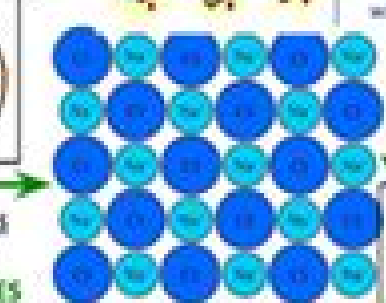
Electrolysis:
 Splitting up a substance using electricity



When there are 2 possible products
 The less reactive forms at the electrode

Electrolysis needs:

- 2 electrodes
- An electrolyte (ionic or molten substance)
- A power supply



Metal ions only arranged in layers

ELECTROSTATIC ATTRACTION between **IONS and ELECTRONS**

Metal ions slide over one another making them **MALLEABLE**

Metals conduct **HEAT** well

NANO-TECHNOLOGY
 Means **REALLY REALLY REALLY** small!!!!

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