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 $\text{br } 1 \text{ n}_2 \text{ 3 f}_2 \text{ 2 nf}_3 \text{ br } 2 \text{ 2 c}_6\text{h}_{10} \text{ 17 o}_2 \text{ 12 co}_2 \text{ 10 h}_2\text{o} \text{ br } 3 \text{ 1 hbr } 1 \text{ khco}_3 \text{ 1 h}_2\text{o} \text{ 1 kbr } 1 \text{ co}_2 \text{ br } 4 \text{ 2 gabr}_3 \text{ 3 na}_2\text{so}_3 \text{ 1 ga}_2 \text{ so}_3 \text{ 3 6 nabr} \text{ br } 5 \text{ 3 sno } 2 \text{ nf}_3 \text{ 3 snf}_2 \text{ 1 n}_2\text{o}_3 \text{ br}$
using the following equation
 $\text{br } 2 \text{ naoh } \text{h}_2\text{so}_4 \text{ 2 h}_2\text{o} \text{ na}_2\text{so}_4 \text{ br}$

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