Question 2

- 2(a)(i) Comparison: X is more easily ionized than Mg.
 - Explanation: X is lower in group II than Mg. The outer electrons of X are in a shell further away from the nucleus than those of Mg. The outer electrons of X are therefore less

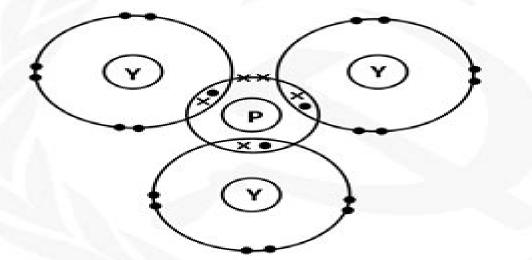
strongly held by the nucleus and more easily lost than the outer electrons of Mg.

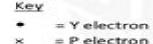
- 2(a)(ii) Comparison: Cl is a stronger oxidizing agent than Y.
 - Explanation: Oxidizing power of elements decreases down group VII. CI is higher in group VII than

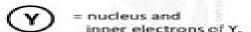
Y. The nucleus of CI is therefore less shielded than that of Y, making external

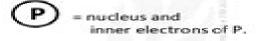
electrons more strongly attracted to the nucleus of CI than of Y.

- 2(b)(i) Both elements are solid at R.T.P.
- 2(b)(ii) At R.T.P. X conducts electricity but Y does not.
- 2(c)(i) Covalent bonding occurs between phosphorous and element Y.
- 2(c)(ii)









2(c)(iii) Chemical formula: PY3-

Two properties of compound: 1. It is a solid with a low melting point.

2. It is soluble in non-polar (organic) solvents.

May June 2014 Practical Chemistry Answer

Rachel Sandford

May June 2014 Practical Chemistry Answer:

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