- 13. Which type of bond is formed when a Lewis acid reacts with a Lewis base?
 - A. Covalent
 - B. Dipole-dipole
 - C. Double
 - D. Hydrogen
- 14. What is the order of increasing acidity of the following acids?

Acid	K,	Acid	pK,
chloroethanoic	1.3×10^{-9}	hydrogen fluoride	3.3
ethanoic	1.7 × 10 ⁻⁶	hydrogen cyanide	9.3

- A. chloroethanoic < ethanoic < hydrogen fluoride < hydrogen cyanide
- B. ethanoic < chloroethanoic < hydrogen fluoride < hydrogen cyanide
- C. chloroethanoic < ethanoic < hydrogen cyanide < hydrogen fluoride
- D. hydrogen cyanide < ethanoic < hydrogen fluoride < chloroethanoic
- 15. Which species behave as Brønsted-Lowry bases in the following reaction?

H₂SO₄ + HNO₃ → H₂NO₃ + HSO₄

- A. HNO₈ and HSO₆-
- B. HNO₃ and H₂NO₃*
- C. H₂SO₄ and HSO₄:
- D. H₂NO₃* and HSO₄*
- 16. Which is a conjugate Brønsted-Lowry acid-base pair?

$$CH_sCOOH(aq) + H_sO(l) \rightleftharpoons CH_sCOO^-(aq) + H_sO^+(aq)$$

- A. CH₃COO- / H₃O-
- B. H₂O / CH₂COO-
- C. H₂O / H₃O ·
- D. CH₃COOH / H₂O
- 17. Which of the following gases does not result in acid deposition?
 - A CON
 - B. NO:
 - C. NO.
 - $D.SO_{B}$
- 18. What occurs when solid sodium hydrogen carbonate reacts with aqueous sulfuric acid?
 - A. Bubbles of sulfur dioxide form.
 - B. Bubbles of both hydrogen and carbon dioxide form.
 - C. Bubbles of hydrogen form.
 - D. Bubbles of carbon dioxide form.
- 19. A student carried out a titration to determine the concentration of an acid and found that his value had good precision but poor accuracy. Which process explains this outcome?
 - A. Consistently overshooting the volume of solution from the burette into the flask.
 - B. Collection of insufficient titration data.
 - C. Reading the meniscus in the burette at a different angle each time.
 - D. Forgetting to rinse the flask after one of the titrations.

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Aubrey Milunsky, Jeff M. Milunsky

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