

Honors Chemistry
Equilibrium Worksheet

Equilibrium: ICE box practice problems

1. In a 10.0L vessel at 1000K, 0.250 mol $\text{SO}_2(\text{g})$ and 0.200 mol $\text{O}_2(\text{g})$ react to form 0.162 mol $\text{SO}_3(\text{g})$ at equilibrium. What is the K_c at 1000K for this reaction?

Balanced equation: $\text{SO}_2(\text{g}) + \text{O}_2(\text{g}) \leftrightarrow \text{SO}_3(\text{g})$

	SO_2	+	O_2	\leftrightarrow	SO_3
I					
C					
E					

2. An aqueous solution of 0.00325M Fe^{3+} was reacted with 0.0020M SCN^- at 25°C. Calculate the value of the equilibrium constant for this reaction when the equilibrium concentration of FeSCN^{2+} was determined to be 0.000667M

	Fe^{3+}	+	SCN^-	\leftrightarrow	FeSCN^{2+}
I					
C					
E					

3. At a certain temperature, 2.0 mole NH_3 is introduced into a 2.0 L container, and the NH_3 partially dissociates into by the following reaction:



At equilibrium, 1.0 mol NH_3 remains, what is the value of K for this reaction?

	$2\text{NH}_3(\text{g})$	\leftrightarrow	$\text{N}_2(\text{g})$	+	$3\text{H}_2(\text{g})$
I					
C					
E					

Ice Box Practice Problems

JS Bruner



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