

Calculations need to be based on the limiting reactant.

- Example 2: What is the mass of sulfur dioxide that can be produced by the reaction of 95.6 g carbon disulfide with 100. g oxygen?



$$\frac{95.6\text{g CS}_2}{76.2\text{g CS}_2} \times \frac{1\text{mol CS}_2}{1\text{mol CS}_2} \times \frac{2\text{mol SO}_2}{1\text{mol CS}_2} \times \frac{64.1\text{g SO}_2}{1\text{mol SO}_2} = 161\text{g SO}_2$$

$$\frac{100.\text{g O}_2}{32.0\text{g O}_2} \times \frac{1\text{mol O}_2}{3\text{mol O}_2} \times \frac{2\text{mol SO}_2}{1\text{mol SO}_2} \times \frac{64.1\text{g SO}_2}{1\text{mol SO}_2} = 134\text{g SO}_2$$

134 g of SO₂ can be produced in this reaction.

Holt Chemistry Limiting Reactants Answers

Kathleen Armour



Holt Chemistry Limiting Reactants Answers:

Adopting the Tune of Term: An Mental Symphony within **Holt Chemistry Limiting Reactants Answers**

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