Explain how you identified the limiting and excess reactant based only on your qualitative observations (not quantitative measurements or calculations). Type your answer into the box. Use complete sentences.

According to observation the initial color of <u>bromogresol</u> purple is yellow in color after adding the <u>NaOH</u>, it turns violet. After Changing the color yellow from purple indicates acid is completely consumed so excess of base addition changes the color into violet.

#### Table C. Calculations to Determine Theoretical Yields

Description	Calculated Value
Volume of NaOH solution	22.20 mL
Initial moles of NaOH	0.002451 moles
Initial moles of H <sub>3</sub> PO <sub>4</sub>	0.002602 moles
Mass of H <sub>2</sub> O produced if NaOH runs out first	0.04414 g
Mass of H <sub>2</sub> O produced if H <sub>3</sub> PO <sub>4</sub> runs out first	1.265 g
Theoretical yield of H <sub>2</sub> O (in grams)	1.265 g
Theoretical yield of Na <sub>3</sub> PO <sub>4</sub> (in grams)	0.4267 g

# Table D. Identification of Limiting and Excess Reactants from Calculations

Description	ption Reactant Name or Formula	
Limiting Reactant	Phosphoric Acid	
Excess Reactant	Sodium Hydroxide	

Explain how you identified the limiting and excess reactant based on your calculations. Be specific. Type your answer into the box. Use complete sentences.

According to our calculation the reacting equivalent of H<sub>3</sub>PO<sub>4</sub> is less than NaOH. Therefore H<sub>2</sub>PO<sub>4</sub> is limiting reactant so product formation depends on H<sub>3</sub>PO<sub>4</sub> because phosphoric acid is completely consumed.

# Formal Lab Report On Limiting Reactant

Joaquín Isac-García, José A.
Dobado, Francisco G. CalvoFlores, Henar Martínez-García

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**ChemDiscovery Lab Manual** Olga I. Agapova, 2002 **Acid Precipitation** ,1985 CHEMISTRY EXPERIMENTS James Signorelli, 2014-09-19 Gifted and talented students and any student interested in pursuing a science major in college needs a rigorous program to prepare them while they are still in high school This book utilizes a format where the application of several disciplines science math and language arts principles are mandated Each lab concludes with either an essay or a detailed analysis of what happened and why it happened This format is based on the expectations of joining a university program or becoming an industrial science professional The ideal student lab report would be written in a lab research notebook and then the essay or final analysis is done on a word processor to allow for repeat editing and corrections The research notebook has all graph pages a title section and a place for the students and their assistants to sign and witness that exercise The basic mechanics of the lab report title purpose procedure diagrams data table math and calculations observations and graphs are handwritten into the book The conclusion is done on a word processor MS Word which allows the instructor to guide the student in writing and editing a complete essay using the MLA format When the final copy is completed the essay is printed and inserted into the lab notebook for grading At the end of the term the student has all their labs in one place for future reference These lab notebooks can be obtained for as little as 3 00 per book This is money well spent In our district the Board of Education buys the books for each student The BOE sees these books as expendable but necessary materials for all science and engineering instruction <u>Improving Student Comprehension in Chemistry</u> <u>Laboratories</u> Tracy Lynn Haroff, 2006 Clinical Laboratory Medicine Kenneth D. McClatchey, 2002 This thoroughly updated Second Edition of Clinical Laboratory Medicine provides the most complete current and clinically oriented information in the field The text features over 70 chapters seven new to this edition including medical laboratory ethics point of care testing bone marrow transplantation and specimen testing providing comprehensive coverage of contemporary laboratory medicine

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Nuclear Science Abstracts ,1975-08 Calculating Theoretical Yield Is Not Impossible! Michael S.

Leonard,2025-02-25 Calculating theoretical and percent yield is a fundamental skill for the laboratory This book primarily targets Organic Chemistry Laboratory courses at the high school or college and university level as a supplemental resource to help students master this skill It begins with simple examples from everyday life demonstrates the importance of balancing the equation addresses the role of the mole in these computations discusses different types of liquids considers the role of significant figures and culminates with the planning of syntheses There are suggestions for further reading as well as practice problems and questions to ensure mastery Begins with examples from everyday life that enable students to understand the concepts of theoretical and percent yield before applying those concepts to the laboratory Addresses the necessity of balancing the reaction equation the centrality of the mole in these calculations and the role of significant figures in reporting the answer Explains how to approach the calculations when using neat liquids or solutions The culmination of this text is the use of the same thought processes to plan the amounts of reactants needed for syntheses of desired quantities of product All of the problems in the book include detailed solutions with accompanying text to explain the answers and ancillaries also include suggestions for further reading

Mine Drainage and Surface Mine Reclamation ,1988

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introduction to all of the significant operations used in the organic lab and includes a large selection of traditional scale and microscale experiments and minilabs Its unique problem solving approach encourages students to think in the laboratory by solving a scientific problem in the process of carrying out each experiment The Second Edition contains a new introductory section Chemistry and the Environment which includes a discussion of the principles of green chemistry Several green experiments have been added and some experiments from the previous editions have been revised to make them greener **ChemCom** American Chemical Society, 1993

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