

A-LEVEL CHEMISTRY

MARK SCHEME

Practice Paper 1

Maximum marks: 105

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1	(a)	(i)	$[H^+][OH^-]$ OR $[H_3O^+][OH^-]$ <i>Ignore (aq)</i> <i>Must have [] not ()</i>	1
		(ii)	$\sqrt{3.46 \times 10^{-14}}$ (= 1.86×10^{-7}) <i>If no square root, CE= 0</i> pH = 6.73 Must be 2dp	2
		(iii)	$[H^+] = 10^{-11.26}$ (= 4.365×10^{-12} OR 4.37×10^{-12}) <i>Mark for working</i>	1
			$K_w = [4.365 \times 10^{-12}$ OR $4.37 \times 10^{-12} \times 0.047] = 2.05 \times 10^{-13}$ <i>Allow $2.05 \times 10^{-13} = 2.1 \times 10^{-13}$</i> <i>Mark for answer</i> <i>Ignore units</i>	1
	(b)	(i)	$HCOOH \rightleftharpoons HCOO^- + H^+$ <i>Must have \rightleftharpoons but ignore brackets.</i> OR $HCOOH + H_2O \rightleftharpoons HCOO^- + H_3O^+$ <i>Allow HCO_2^- or $CHOO^-$ ie minus must be on oxygen, so penalise $COOH^-$</i>	1
		(ii)	$K_a = \frac{[H^+][HCOO^-]}{[HCOOH]}$ OR $\frac{[H_3O^+][HCOO^-]}{[HCOOH]}$ <i>Must have all brackets but allow ()</i> <i>Must be HCOOH etc.</i> <i>Allow ecf in formulae from (b)(i)</i>	1
		(iii)	M1 $K_a = \frac{[H^+]^2}{[HCOOH]}$ ($[H^+]^2 = 1.78 \times 10^{-4} \times 0.056 = 9.97 \times 10^{-6}$) <i>Allow HA or HX etc.</i> <i>Allow $[H^+] = \sqrt{K_a \times [HA]}$ for M1</i>	1
			M2 $[H^+] = 3.16 \times 10^{-3}$ <i>Mark for answer</i>	1
			M3 pH = 2.50 allow more than 2 dp but not fewer <i>Allow correct pH from their wrong $[H^+]$ here only if square root shown but not taken, pH = 5.00 can score max 2 for M1 and M3</i>	1

May 200a Level Chemistry Paper

Tobias Bleicker



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