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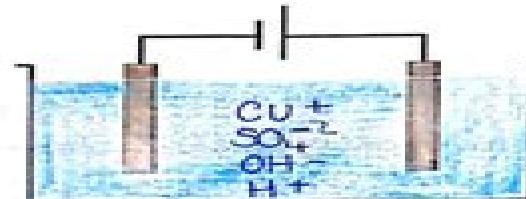
10. Electrolysis of copper – refining copper

Electrodes: metallic copper

Electrolyte: Copper sulfate solution

Ions Present: Cu^{2+} , SO_4^{2-} , OH^- , H^+

Half equations:



Observations:

Cathode: Dark solid deposit (pure copper)

Anode: Desintegration of anode and yellow deposit below

Uses of copper: electrical wiring, copper plating



(show in the graph how it will look like)

Why is it necessary to purify Copper? In this way, it will have more conductivity

What are the elements present in the slag? Ag , Au , Pt

11. Electropatting with Copper

Electrodes:

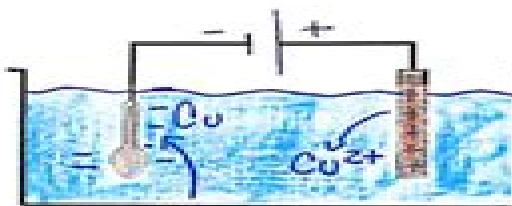
Cathode: Spoon (metal more reactive than Cu)

Anode: Pure copper

Electrolyte: Copper Sulfate or Nitrate

Ions Present: Cu^{2+} , SO_4^{2-} , OH^- , H^+

Half equations:



Observations:

Cathode: the color of the spoon becomes darker

Anode: Disintegrate Copper

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