

### Question 1 of 10

Ionization energy increases down a group because the number of protons increases and it's more difficult to lose electrons because they are being pulled closer.

☐ False

☐ True

### Question 2 of 10

Which of the following has the largest ionization energy?

☐ N

☐ Si

☐ Al

☐ P

### Question 3 of 10

Which of the following elements has the lowest ionization energy?

☐ Cesium (Cs)

☐ Sodium (Na)

☐ Lithium (Li)

☐ Rubidium (Rb)

### Question 4 of 10

Order the elements in decreasing ionization energy. Na, S, Al, F, K, O

☐ Al, O, K, F, S, Na

☐ F, O, S, Al, Na, K

Na, S, F, K, O, Al

# Ionization Energy Exams Study Guide Answers

**M Woodhall**



## **Ionization Energy Exams Study Guide Answers:**

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